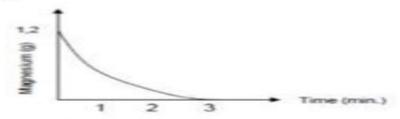
# Grade 11 Physical Science M...

records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

The change in the mass of magnesium during the reaction is shown on the graph below.



- 6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer.
- 6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes.

# **Hide Answers**

#### QUESTION/VRAAG 6

- 6.1.1 2H<sub>2</sub>O<sub>2</sub> (aq) ✓ → 2H<sub>2</sub>O (ℓ) + O<sub>2</sub> (g) ✓
- 6.1.2 The catalyst lowers the activation energy of the reaction 
  Accept: catalyst speeds up the reaction
  In Katalisator vertiage die aktiveringsenergie van die reaksie
  - 'n Katalisator verlaag die aktiveringsenergie van die reaksie 🗸 🗸 Aanvaar: katalisator laat die reaksie vinniger plaasvind
- 6.1.3 OPTION 1/OPSIE 1

$$0,049 = \frac{m}{34} \checkmark$$

OPTION 2/OPSIE 2
From the balanced equation:
Vanaf gebalanseerde
vergelyking:

68g  $H_2O_2 \rightarrow 24,45 \text{ dm}^3 O_2 \checkmark \checkmark$   $X \text{ g } H_2O_2 \rightarrow 600 \times 10^{-3} \text{ dm}^3 \checkmark$  $X = 68 \times 0.6 \checkmark$  (2)

(7) [19]

(2)

$$X = \frac{68 \times 0.6}{24.45}$$
   
  $X = 1.67 g$ 

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