Chemical Bonding

SECTION 1

Name.

SHORT ANSWER	Answer the following questions in the space provided.

- a ___ A chemical bond between atoms results from the attraction between the valence electrons and _____ of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures
- A covalent bond consists of (a) a shared electron. (c) two different ions. (b) a shared electron pair. (d) an octet of electrons.
- a If two covalently bonded atoms are identical, the bond is identified as (c) ionic. (a) nonpolar covalent. (b) polar covalent. (d) dipolar.
- A covalent bond in which there is an unequal attraction for the shared electrons is (a) nonpolar. (c) ionic.
 - (b) polar. (d) dipolar. Atoms with a strong attraction for electrons they share with another atom
 - exhibit (a) zero electronegativity. (c) high electronegativity.
 - (b) low electronegativity. (d) Lewis electronegativity.
- Bonds that possess between 5% and 50% ionic character are considered to be
 - (a) ionic. (c) polar covalent. (b) pure covalent. (d) nonpolar covalent.
- The greater the electronegativity difference between two atoms bonded together, the greater the bond's percentage of
 - (a) ionic character. (c) metallic character. (b) nonpolar character. (d) electron sharing.
- 8. The electrons involved in the formation of a chemical bond are called valence electrons
- A chemical bond that results from the electrostatic attraction between positive and negative ions is called a(n) _____ionic bond

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Questions And Answers On Chemical Bonding

Stanley E. Manahan

Questions And Answers On Chemical Bonding:

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