

CHEMISTRY



Periodic Table of Elements

Properties of the Periodic Table

- The periodic table is a tabular arrangement of the chemical elements, organized on the basis of their atomic number, electron configuration, and recurring chemical properties.
- The elements are arranged in rows (periods) and columns (groups).
- The periodic table is a fundamental tool in chemistry, used to predict the properties of elements and to understand the relationships between them.

Atomic Trends

- Atomic size increases from top to bottom and from right to left.
- Ionization energy increases from bottom to top and from left to right.
- Electronegativity increases from bottom to top and from left to right.

Atomic Structure

Atoms are made up of three subatomic particles: electrons, protons, and neutrons. Electrons are negatively charged, protons are positively charged, and neutrons are neutral.

The atomic number (Z) is the number of protons in the nucleus of an atom. The mass number (A) is the sum of the number of protons and neutrons in the nucleus.

Atomic Structure: Protons & Electrons

Protons and electrons are present in all atoms. The number of protons is equal to the number of electrons, making the atom electrically neutral.

The mass of a proton is approximately 1836 times the mass of an electron.

Properties of Matter

Property	Solid	Liquid	Gas
Shape	Definite	Indefinite	Indefinite
Volume	Definite	Definite	Indefinite
Compressibility	Low	Medium	High
Diffusibility	Low	Medium	High

States of Matter

Matter exists in three states: solid, liquid, and gas. The state of matter depends on the temperature and pressure.

Changes in state are called phase transitions. Examples include melting, boiling, and condensation.

Atomic Structure: Neutrons

Neutrons are found in the nucleus of an atom. They are neutral particles with a mass similar to that of a proton.

The number of neutrons in an atom is given by the mass number minus the atomic number.

Atomic Structure: Isotopes

Isotopes are atoms of the same element that have the same number of protons but different numbers of neutrons.

Isotopes have the same chemical properties but different physical properties.

Atomic Structure: Radioactivity

Radioactivity is the spontaneous emission of particles or energy from an unstable nucleus.

There are three types of radioactive decay: alpha, beta, and gamma.

Atomic Structure: Applications

Radioactive isotopes are used in various fields, including medicine, agriculture, and industry.

For example, carbon-14 is used in radiocarbon dating to determine the age of ancient artifacts.

Atomic Structure: Molecules

Molecules are formed by the combination of two or more atoms. They can be made of the same element (e.g., O₂) or different elements (e.g., H₂O).

The molecular formula represents the number and kind of atoms in a molecule.

Atomic Structure: Compounds

Compounds are substances formed by the chemical combination of two or more elements in a fixed ratio.

The chemical formula represents the composition of a compound.

Atomic Structure: Mixtures

Mixtures are combinations of two or more substances that are not chemically combined.

Mixtures can be separated into their constituent parts by physical means.

Atomic Structure: Solutions

Solutions are homogeneous mixtures of two or more substances. The substance being dissolved is the solute, and the substance doing the dissolving is the solvent.

Solutions can be formed in solid, liquid, or gas states.

Atomic Structure: Crystals

Crystals are solid materials whose atoms, molecules, or ions are arranged in a highly ordered, repeating pattern.

Crystals have characteristic shapes and properties.

Atomic Structure: Polymers

Polymers are long chains of repeating units called monomers. They can be natural (e.g., rubber) or synthetic (e.g., plastic).

Polymers have a wide range of properties and uses.

Atomic Structure: Nanotechnology

Nanotechnology is the study and manipulation of matter at the nanoscale, which is between 1 and 100 nanometers.

Nanotechnology has many applications in medicine, electronics, and materials science.

Atomic Structure: Environmental Chemistry

Environmental chemistry is the study of chemical processes in the environment. It deals with the behavior of chemicals in the atmosphere, water, and soil.

Understanding environmental chemistry is crucial for addressing global issues like climate change and pollution.

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