

## Quick Lab

**Purpose** To use a graph to identify period and group trends

### Materials

- graph paper
- pencil



## Periodic Trends in Ionic Radii

### Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

### Analyze and Conclude

- 1. Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- 2. Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- 3. Describe** What is the shape of a portion of the graph that corresponds to one period?
- 4. Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- 5. Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

**Figure 6.23 Atomic and Ionic Radii**  
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



# Quick Lab Periodic Trends In Ionic Radii Key Answers

**DP Hallahan**



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