

## Blackline Master 5.8

# Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

(a) magnesium oxide	<u>MgO</u>	(k) copper(I) bromide	<u>CuBr</u>
(b) sodium fluoride	<u>NaF</u>	(l) tin(II) iodide	<u>SnI<sub>2</sub></u>
(c) aluminum nitride	<u>AlN</u>	(m) iron(III) chloride	<u>FeCl<sub>3</sub></u>
(d) potassium sulfide	<u>K<sub>2</sub>S</u>	(n) calcium phosphide	<u>Ca<sub>3</sub>P<sub>2</sub></u>
(e) lithium iodide	<u>LiI</u>	(o) lead(II) oxide	<u>PbO</u>
(f) calcium bromide	<u>CaBr<sub>2</sub></u>	(p) lead(IV) fluoride	<u>PbF<sub>4</sub></u>
(g) beryllium oxide	<u>BeO</u>	(q) tin(IV) bromide	<u>SnBr<sub>4</sub></u>
(h) nickel <sup>(II)</sup> chloride	<u>NiCl<sub>2</sub></u>	(r) copper(II) sulfide	<u>CuS</u>
(i) magnesium nitride	<u>Mg<sub>3</sub>N<sub>2</sub></u>	(s) iron(II) oxide	<u>FeO</u>
(j) aluminum sulfide	<u>Al<sub>2</sub>S<sub>3</sub></u>	(t) calcium nitride	<u>Ca<sub>3</sub>N<sub>2</sub></u>

2. Write the names for the following compounds.

(a) Li <sub>2</sub> O	<u>lithium oxide</u>	(k) PbS	<u>lead (II) sulfide</u>
(b) AlCl <sub>3</sub>	<u>aluminum chloride</u>	(l) SnO <sub>2</sub>	<u>tin (IV) oxide</u>
(c) MgS	<u>magnesium sulfide</u>	(m) Na <sub>2</sub> S	<u>sodium sulfide</u>
(d) CaO	<u>calcium oxide</u>	(n) Mg <sub>3</sub> P <sub>2</sub>	<u>magnesium phosphide</u>
(e) KBr	<u>potassium bromide</u>	(o) NiO	<u>nickel (II) oxide</u>
(f) BeF	<u>beryllium fluoride</u>	(p) CuI	<u>copper (I) iodide</u>
(g) Na <sub>3</sub> N	<u>sodium nitride</u>	(q) PbCl <sub>2</sub>	<u>lead (II) chloride</u>
(h) Al <sub>2</sub> O <sub>3</sub>	<u>aluminum oxide</u>	(r) FeP	<u>iron (III) phosphide</u>
(i) CuCl <sub>2</sub>	<u>copper (II) chloride</u>	(s) CaF <sub>2</sub>	<u>calcium fluoride</u>
(j) FeBr <sub>3</sub>	<u>iron (III) bromide</u>	(t) K <sub>3</sub> P	<u>potassium phosphide</u>

# Physical Science Names And Formulas Worksheet 1

**Chandan Sukumar Sengupta**



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