

Mathematics

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(Chapter - 3) (Pair of Linear Equations in two variables) (Exercise 3.3)
(Class X)

Question 1:

Solve the following pair of linear equations by the elimination method and the substitution method:

(i) $x + y = 5$ and $2x - 3y = 4$

(ii) $3x + 4y = 10$ and $2x - 2y = 2$

(iii) $3x - 5y - 4 = 0$ and $9x = 2y + 7$

(iv) $\frac{x}{2} + \frac{2y}{3} = -1$ and $x - \frac{y}{3} = 3$

Answer 1:

(i) $x + y = 5$

--- (1)

$2x - 3y = 4$

--- (2)

Multiply equation (1) by 2 and subtracting from equation (2), we get

$$2x - 3y = 4$$

$$2x + 2y = 10$$

$$\hline -5y = -6$$

$$\Rightarrow y = \frac{6}{5}$$

Putting the value of y in equation (1), we get

$$x + \frac{6}{5} = 5 \Rightarrow x = 5 - \frac{6}{5} = \frac{19}{5} \Rightarrow x = \frac{19}{5}$$

Hence, $x = \frac{19}{5}$ and $y = \frac{6}{5}$

(ii) $3x + 4y = 10$

--- (1)

$2x - 2y = 2$

--- (2)

Multiply equation (2) by 2 and adding with equation (1), we get

$$3x + 4y = 10$$

$$4x - 4y = 4$$

$$\hline 7x = 14$$

$$\Rightarrow x = \frac{14}{7} = 2$$

Putting the value of x in equation (1), we get

$$3(2) + 4y = 10 \Rightarrow 4y = 4 \Rightarrow y = 1$$

Hence, $x = 2$ and $y = 1$.

(iii) $3x - 5y - 4 = 0$

--- (1)

$9x - 2y = 7$

--- (2)

Multiply equation (1) by 3 and subtracting from equation (2), we get

$$9x - 2y = 7$$

$$9x - 15y = 12$$

$$\hline -13y = -5$$

$$\Rightarrow y = -\frac{5}{13}$$

Putting the value of y in equation (2), we get

$$9x - 2\left(-\frac{5}{13}\right) = 7$$

$$\Rightarrow 9x = 7 - \frac{10}{13} = \frac{81}{13} \Rightarrow x = \frac{9}{13}$$

Hence, $x = \frac{9}{13}$ and $y = -\frac{5}{13}$

(iv) $\frac{x}{2} + \frac{2y}{3} = -1$

$$\Rightarrow 3x + 4y = -6$$

--- (1)

and $x - \frac{y}{3} = 3$

$$\Rightarrow 3x - y = 9$$

--- (2)

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+ \text{aq} + 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + \text{SO}_4^{2-} \text{aq} \rightarrow 2\text{Na}^+ \text{aq} + \text{SO}_4^{2-} \text{aq} + \text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{Na}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{Na}^+ \text{aq} + \text{Cl}^- \text{aq} + \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ the spectator ions are Cl^- and Na^+ net ionic equation $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$ here is a video on spectator ions

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ net ionic $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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web the $\text{ag}^+ \text{aq}$ and $\text{cl}^- \text{aq}$ ions become agcl s but the $\text{na}^+ \text{aq}$ ions and the $\text{no}_3^- \text{aq}$ ions stay as $\text{na}^+ \text{aq}$ ions and $\text{no}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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