

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Moles of product formed from a limiting reagent using stoichiometry. Percent yield = (actual yield / theoretical yield) x 100.

Chemical reactions involve the breaking and forming of chemical bonds. In a chemical reaction, the atoms of the reactants are rearranged to form the products.



The reaction of hydrogen and oxygen to form water is exothermic.

2H ₂ (g)	20.0 g
O ₂ (g)	16.0 g
H ₂ O(l)	36.0 g
H ₂ (g)	0.0 g

1. 20.0 g of hydrogen and 16.0 g of oxygen react to form 36.0 g of water. What is the limiting reagent? What is the percent yield?



2. 20.0 g of H₂(g) is the limiting reagent. 16.0 g of O₂(g) is the excess reagent.

3. 36.0 g of H₂O(l) is the theoretical yield. 36.0 g of H₂O(l) is the actual yield.



4. 36.0 g of H₂O(l) is the theoretical yield. 36.0 g of H₂O(l) is the actual yield. 100% yield.



Percent Yield And Limiting Reactants Quiz

Paul E. Curtis, Roger L. Courson

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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