

Reaction Rates and Potential Energy Diagrams

1. Chemical reactions occur when reactants collide. For what reasons may a collision fail to produce a chemical reaction?

Not enough energy; improper angle.

2. If every collision between reactants leads to a reaction, what determines the rate at which the reaction occurs?

Nature of reactants, Concentration, Temperature, Catalysts.

3. What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

E_a is the minimum amount of energy for a reaction to occur. It is the amount of energy required to create an activated complex.

4. What happens when a catalyst is used in a reaction?

A catalyst changes the reaction mechanism, in the process lowering the activation energy.

5. Name 4 things that will speed up or slow down a chemical reaction.

Increase concentration by distillation of a solvent, Increase concentration by increasing pressure of a gas, Increase temp, Add a catalyst, Add an inhibitor.

6. Draw an energy diagram for a reaction. Label the axis, PE of reactants = 350 KJ/mol, E_a = 100 KJ/mol, PE of products = 250 KJ/mol.



7. Is the reaction in # 6 exothermic or endothermic? Explain.

Exothermic. The ΔH is -100 KJ/mol which means heat is released.

8. How could you lower the activation energy for the reaction in #6?

Add a catalyst.

Potential Energy Diagram Practice Problems

David R. Klein



Potential Energy Diagram Practice Problems:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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