Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale) 2.1 1.0	
H	31	1312		
I., i	128	520		
Nia	166	496	0.9	
K	203	410	0.8	
Rb	220	403	0.8	
Cs	244	376	0.7	

Symbol	Atomic Radius (Picometers)	First lonization energy (kilojoules/mole)	Electronegativity (4-point scale) 0.9 1.2	
Na	166	496		
Mg	141	738		
AL	121	578	1.5	
Sii	111	787	1.8	
P	107	1012	2.1	
S	105	1000	2.5	
CL	102	1251	3.0	
Ar	106	1521		

- Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
- Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First
  ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are
  available, as long as you are consistent.
- 3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
- Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

Li	Across	a period (-	→) electrone	gativity ter	nds to		
Na 166 496 0.9	Mg	Al	Si	Р	s	СІ	A
к	Down	a Group (	atomic rac	lius tends t	o		
Rb	0.0000000000000000000000000000000000000		first ioniza		tends to _		

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#### Table of Contents Periodic Table Trends Worksheet Answers Bisd303

- 1. Understanding the eBook Periodic Table Trends Worksheet Answers Bisd303
  - The Rise of Digital Reading Periodic Table Trends Worksheet Answers Bisd303
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Table Trends Worksheet Answers Bisd303
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Trends Worksheet Answers Bisd303
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Table Trends Worksheet Answers Bisd303
  - Personalized Recommendations
  - Periodic Table Trends Worksheet Answers Bisd303 User Reviews and Ratings
  - Periodic Table Trends Worksheet Answers Bisd303 and Bestseller Lists

- 5. Accessing Periodic Table Trends Worksheet Answers Bisd303 Free and Paid eBooks
  - Periodic Table Trends Worksheet Answers Bisd303 Public Domain eBooks
  - Periodic Table Trends Worksheet Answers Bisd303 eBook Subscription Services
  - Periodic Table Trends Worksheet Answers Bisd303 Budget-Friendly Options
- 6. Navigating Periodic Table Trends Worksheet Answers Bisd303 eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Periodic Table Trends Worksheet Answers Bisd303 Compatibility with Devices
  - Periodic Table Trends Worksheet Answers Bisd303 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Table Trends Worksheet Answers Bisd303
  - Highlighting and Note-Taking Periodic Table Trends Worksheet Answers Bisd303
  - Interactive Elements Periodic Table Trends Worksheet Answers Bisd303
- 8. Staying Engaged with Periodic Table Trends Worksheet Answers Bisd303
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Table Trends Worksheet Answers Bisd303
- 9. Balancing eBooks and Physical Books Periodic Table Trends Worksheet Answers Bisd303
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Periodic Table Trends Worksheet Answers Bisd303
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Trends Worksheet Answers Bisd303
  - Setting Reading Goals Periodic Table Trends Worksheet Answers Bisd303
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Trends Worksheet Answers Bisd303
  - Fact-Checking eBook Content of Periodic Table Trends Worksheet Answers Bisd303
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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