Using the Periodic Table

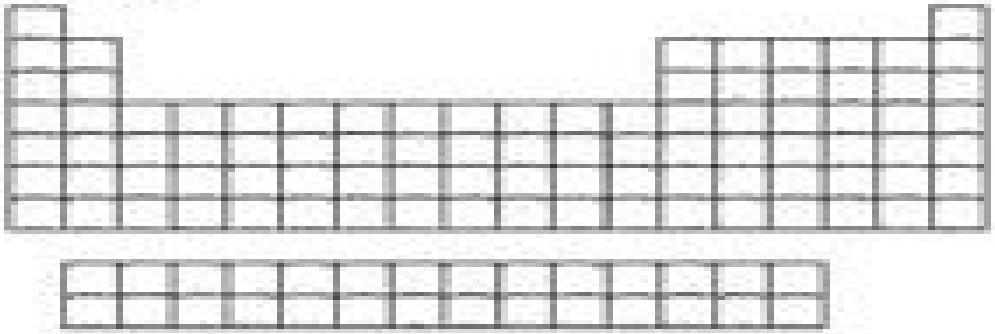
C - Introduction to the Periodic Table

6) a. Using the blank Periodic Table provided & your understanding of atomic number, describe how

the elements are listed on the Periodic table.

b. Using the Periodic Table to the right, draw a circle around a "group", and fill in one "period".

c. Based on the evidence we have gathered from scientific research, what do we know about elements that are found in the same group on the Periodic Table?



7) Give the symbol and atomic number for each of the following elements and indicate whether each element is a metal, or a nonmetal, and whether they are members of a named group/family on the Periodic Table.

Name	Symbol	Atomic Number	Metal or Nonmetal	Group/Family Name
Argon				
Chlorine				
Barium				
Calcium				

8) Based on your knowledge of the Periodic Table and the natural states that elements occur in, provide an example of:

a. An element referred to as a noble gas –

b. An element that naturally occurs as a diatomic molecule –

c. An element that occurs as a solid at room temperature –

d. An element that occurs as a liquid at room temperature –

e. An element that occurs as a gas at room temperature –

Ions & Their Compounds

9) Why does an atom have a zero net charge?

10) Ions are formed by an atom gaining or losing _____

11) Metals naturally form what type of ion? Nonmetals? Provide definitions for each type of ion.

12) Describe what happens to

a. the size of an atom when it loses an electron

b. the size of an atom when it gains an electron

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Arshad Iqbal

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