

**Physical Science – 2nd Semester – Final Exam Study Guide – (2012-2013)****1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- i. Mass of an object? triple beam balance Basic SI Unit? kg
  - ii. Volume of an object? ruler or water displacement method Basic SI Unit? cm<sup>3</sup> or ml
  - iii. Length of an object? Ruler Basic SI Unit? meter
  - iv. Density of an object? graduated cylinder Basic SI Unit? ml
- b. What is the formula for calculating density? mass/volume
- i. What is the density of a cube that is 15 g and has a volume of 3 cm<sup>3</sup>? 5g/cm<sup>3</sup>
  - ii. Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has.  
solid H<sub>2</sub>O is less dense than liquid H<sub>2</sub>O

**2. Chemistry Basics**

- a. Element (def.) substance made up of one kind of atom
- b. Atom (def.) basic building block of matter
- c. Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- d. Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

**3. Model of the Atom: Evolution of the Atomic Theory**

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

| Scientist         | Year        | Major Discovery/ Contribution                                                                                                                                                                                          | Name of atomic model       |
|-------------------|-------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------|
| John Dalton       | <u>1802</u> | Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.                                                                                               | <u>Billiard Ball Model</u> |
| J.J. Thomson      | <u>1897</u> | Negatively charged electron                                                                                                                                                                                            | <u>Plum Pudding Model</u>  |
| Niels Bohr        | <u>1913</u> | Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> ; Atoms absorb or give off energy when the electrons move from one shell ( <u>energy level</u> ) to another | <u>Bohr Model</u>          |
| Ernest Rutherford | <u>1909</u> | 1909 ( <u>Based on idea that positive reflects positive</u> ) atom mostly empty space; electrons orbit randomly around positively charged nucleus                                                                      | <u>Planetary Model</u>     |

**4. Periodic Table**

- a. The current periodic table is arranged by increasing atomic number.
- b. **Reading the Periodic Table**
- i. An element's properties can be predicted from its location in the periodic table.
  - ii. The atomic number represents the number of protons.
  - iii. The atomic mass represents the number of neutrons and protons.

# Physical Science 2nd Semester Study Guide 2012 2013

**JA Banks**



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*Research in Education* ,1973-12

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## **Table of Contents Physical Science 2nd Semester Study Guide 2012 2013**

1. Understanding the eBook Physical Science 2nd Semester Study Guide 2012 2013
  - The Rise of Digital Reading Physical Science 2nd Semester Study Guide 2012 2013
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science 2nd Semester Study Guide 2012 2013
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science 2nd Semester Study Guide 2012 2013
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science 2nd Semester Study Guide 2012 2013
  - Personalized Recommendations
  - Physical Science 2nd Semester Study Guide 2012 2013 User Reviews and Ratings
  - Physical Science 2nd Semester Study Guide 2012 2013 and Bestseller Lists
5. Accessing Physical Science 2nd Semester Study Guide 2012 2013 Free and Paid eBooks
  - Physical Science 2nd Semester Study Guide 2012 2013 Public Domain eBooks
  - Physical Science 2nd Semester Study Guide 2012 2013 eBook Subscription Services
  - Physical Science 2nd Semester Study Guide 2012 2013 Budget-Friendly Options

6. Navigating Physical Science 2nd Semester Study Guide 2012 2013 eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science 2nd Semester Study Guide 2012 2013 Compatibility with Devices
  - Physical Science 2nd Semester Study Guide 2012 2013 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science 2nd Semester Study Guide 2012 2013
  - Highlighting and Note-Taking Physical Science 2nd Semester Study Guide 2012 2013
  - Interactive Elements Physical Science 2nd Semester Study Guide 2012 2013
8. Staying Engaged with Physical Science 2nd Semester Study Guide 2012 2013
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science 2nd Semester Study Guide 2012 2013
9. Balancing eBooks and Physical Books Physical Science 2nd Semester Study Guide 2012 2013
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science 2nd Semester Study Guide 2012 2013
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science 2nd Semester Study Guide 2012 2013
  - Setting Reading Goals Physical Science 2nd Semester Study Guide 2012 2013
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science 2nd Semester Study Guide 2012 2013
  - Fact-Checking eBook Content of Physical Science 2nd Semester Study Guide 2012 2013
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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