

Part A

Charles's law states that the volume (V) of a fixed quantity of gas is directly proportional to its temperature (T) at a constant pressure.

$$V \propto T$$

You can verify this law by plotting the graph of a gas's volume versus its temperature.

To perform this analysis, pump the handle only once so that a fixed number of gas molecules enter the gas chamber. Set the number of "Heavy Species" gas molecules to 100 using the text box given in the tab named **Gas in chamber**. Once the pressure reaches the value of about 0.50 atm, click on the "Pressure" button under the tab **Constant Parameter**, which is at the top right corner of the simulation. Go to the panel named "Tools and Options." Select the ruler by checking off the option in the Measurement Tools. Observe that the height of the cylinder (as measured left to right) does not remain constant because the molecules exert pressure on the walls of the cylinder.

Set the temperature by using the heat control box to add or remove heat as given in the table below.

Temperature (K)	200.	250.	300.	350.
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Use the ruler to find the height of the cylinder as an average between two whole nanometer markings on the ruler. For example, if the value is fluctuating between 6.0 and 7.0 nm, consider 6.5 nm to be the height of the cylinder even if the ruler appears to hover closer to either marking.

Complete the table below with your raw data for the height of the cylinder at each temperature.

Drag the appropriate labels to their respective targets.

3.5

4.5

5.5

6.5

7.5

8.5

Temperature (K)	200.	250.	300.	350.
Height of cylinder (nm)	4.5	5.5	6.5	7.5

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Phet Gas Properties Answers

Adolf R. Hochstim



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The Stock Exchange Gazette, 1952 **Western Industry**, 1956 Gas Properties, Ideal Gas Properties Jing Chao, 1986 Gas Properties, Gas Properties, Present State of Our Knowledge of Gas Properties, 1947 COMPENDIUM OF GAS PROPERTIES A-6C3 Fluids Committee, 2008

Experimental Determination of Gas Properties at High Temperature And/or Pressures C. Carey, 1974 *Gas Measurement Manual: Measurement of gas properties* American Gas Association, 1977 **Gas Properties and Compressor Data**, 1981 **Gas Properties** Lifeliqe, 2019 This lesson plan covers gas compression gas pressure and using a vapor pressure curve to determine boiling points at different atmospheric pressures Gas Measurement Manual American Gas Association (AGA), 1988 *Experimental Determination of Gas Properties at High Temperatures And/or Pressures* C. Carey, E. H. Carnevale, Arnold Engineering Development Center, S. Uva, T. Marshall, PANAMETRICS INC WALTHAM MASS., 1968 The problem of extending transport property measurements to conditions of high pressure p

Teaching the Gas Properties and Gas Laws Michael T. Hammar, 2013 As a chemistry teacher I felt that the gas laws were rather simple concepts to understand My students however felt the opposite Conceptual issues as well as mathematical issues were problematic every year In this study a unit about gas properties and gas laws was modified to include inquiry based teaching methods The research questions focused on how these changes affected student results on a traditional end of the unit test and on an alternative assessment Students attitudes during these lessons were also monitored to see if the inquiry methods improved student perceptions of teaching learning and self awareness The results of this study showed that an inquiry approach improved the student's ability to perform on a traditional end of the unit test in the areas of microscopic understanding atomic level symbolic understanding mathematical level and graphical understanding relationships between pressure volume and temperature It was also shown to have a positive effect on the student's confidence and attitude towards the material presented when inquiry was utilized On the other hand the revised unit did not improve the student's overall performance on an alternative assessment geared towards the concept of pressure but showed some improvement in two specific areas a student's understanding that pressure is created by molecules and their movement and b student's understanding that when pressure is equalized little happens within a system

Gas Properties Computational Procedure Suitable for Electronic Calculators James Randolph Andrews, Oscar Biblarz, NAVAL POSTGRADUATE SCHOOL MONTEREY

CALIF., Naval Postgraduate School (U.S.), 1974 A calculating procedure is presented based on a least squares polynomial approximation for duplicating existing tabular values of selected thermodynamic and transport properties of various important gases Most of the data refer to ideal gases however for functions of both temperature and pressure the pressure has been specified such that the polynomial approximations have only temperature as the independent variable Suggested algorithms for polynomial evaluations up to the sixth degree are presented and optimized for hand held or desk top calculators Using the suggested polynomial fits and a suitable calculator it is possible to duplicate existing table values for the various functions of the selected gases without interpolation of reference to the tables per se The error of the included fits is generally less than 0.5% With the lowest order approximation the error can be as high as 1% Author **Combustion Gas Properties**, 1985 **Combustion Gas Properties**, 1985 **Gas Properties Behind Shocks at Hypersonic Velocities** Adolf R. Hochstim, 1957

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