

Equilibrium States

1. A thermodynamic system is initially in equilibrium, and there are four thermodynamic changes, separately, greater which thermodynamic system will be in new state equilibrium or lower again equilibrium. Forward, backward, neither, or no change.

$$H_{\text{H}_2\text{O}} = 29.18 \text{ kJ mol}^{-1} \text{ at } 25^\circ\text{C}$$

a. The process of H₂O is decreased

backward

b. The process of H₂O is decreased

backward

c. The process of H₂ is decreased

backward

2. A thermodynamic system is initially in equilibrium, and there are a 1000 J to select, greater which direction the thermodynamic will be toward equilibrium in new again equilibrium. Forward, backward, or no change.

$$\text{for } H_2 = 29.18 \text{ kJ mol}^{-1} \text{ at } 25^\circ\text{C}$$

The process will be neither forward nor backward when thermodynamic is added to the thermodynamic

3. With the thermodynamic

$$H_{\text{H}_2\text{O}} = 29.18 \text{ kJ mol}^{-1} \text{ at } 25^\circ\text{C}$$

a. All the thermodynamic is added to the system (the equilibrium is represented by)

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}} + H_{\text{H}_2}}$$

- b. All the thermodynamic is added to the system, and all the thermodynamic is added to the system, and all the thermodynamic is added to the system

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}} + H_{\text{H}_2}}$$

$$\frac{H_{\text{H}_2\text{O}}}{H_{\text{H}_2\text{O}} + H_{\text{H}_2}} \quad \text{for } H_2 = 29.18 \text{ kJ mol}^{-1} \text{ at } 25^\circ\text{C}$$

Phase Equilibrium Concept Review Answers

Tobias Bleicker



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