

The letters "pH" represent the French words "pouvoir hydrogene" which means "hydrogen power".

- The definition of pH is *pH is equal to the negative log (logarithm) of the hydrogen ion concentration of a solution.*
- The logarithm of a number is the power to which 10 must be raised to equal that number.

A pH value of less than 7 indicates a(n) ACIDIC solution. A pH value of 7 indicates a neutral solution. A pH value of more than 7 indicates a(n) BASIC solution.

PROBLEMS: Show all work and circle the final answer.

1. Determine the pH of a 0.010 M HNO_3 solution.

$$\begin{aligned} \text{pH} &= -\log[0.010] \\ &= \boxed{2} \end{aligned}$$

2. What is the pH of a 2.5×10^{-6} M solution of HCl?

$$\begin{aligned} \text{pH} &= -\log[2.5 \times 10^{-6}] \\ &= \boxed{5.6} \end{aligned}$$

3. Calculate the pH of a solution of 0.0025M H_2SO_4 .

$$\begin{aligned} \text{pH} &= -\log[2 \times 0.0025] \\ &= -\log[0.005] \\ &= \boxed{2.3} \end{aligned}$$

4. Calculate the pH of a 0.0010 M NaOH solution.

$$\begin{aligned} [\text{OH}^-][0.0010] &= 1.0 \times 10^{-14} \\ [\text{OH}^-] &= \frac{1.0 \times 10^{-14}}{0.0010} \\ \text{pH} &= -\log[1 \times 10^{-11}] \\ &= \boxed{11} \end{aligned}$$

5. What is the pH of a 0.020M $\text{Sr}(\text{OH})_2$ solution?

$$\begin{aligned} [\text{OH}^-][0.04] &= 1 \times 10^{-14} \\ [\text{OH}^-] &= 2.5 \times 10^{-2} \\ \text{pH} &= -\log[4] \\ &= -\log 2.5 \times 10^{-2} \end{aligned}$$

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1. Understanding the eBook Ph Calculations Answers With Work Chemfiesta
 - The Rise of Digital Reading Ph Calculations Answers With Work Chemfiesta
 - Advantages of eBooks Over Traditional Books
2. Identifying Ph Calculations Answers With Work Chemfiesta
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph Calculations Answers With Work Chemfiesta
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Calculations Answers With Work Chemfiesta
 - Personalized Recommendations
 - Ph Calculations Answers With Work Chemfiesta User Reviews and Ratings

- Ph Calculations Answers With Work Chemfiesta and Bestseller Lists
- 5. Accessing Ph Calculations Answers With Work Chemfiesta Free and Paid eBooks
 - Ph Calculations Answers With Work Chemfiesta Public Domain eBooks
 - Ph Calculations Answers With Work Chemfiesta eBook Subscription Services
 - Ph Calculations Answers With Work Chemfiesta Budget-Friendly Options
- 6. Navigating Ph Calculations Answers With Work Chemfiesta eBook Formats
 - ePub, PDF, MOBI, and More
 - Ph Calculations Answers With Work Chemfiesta Compatibility with Devices
 - Ph Calculations Answers With Work Chemfiesta Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ph Calculations Answers With Work Chemfiesta
 - Highlighting and Note-Taking Ph Calculations Answers With Work Chemfiesta
 - Interactive Elements Ph Calculations Answers With Work Chemfiesta
- 8. Staying Engaged with Ph Calculations Answers With Work Chemfiesta
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph Calculations Answers With Work Chemfiesta
- 9. Balancing eBooks and Physical Books Ph Calculations Answers With Work Chemfiesta
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph Calculations Answers With Work Chemfiesta
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ph Calculations Answers With Work Chemfiesta
 - Setting Reading Goals Ph Calculations Answers With Work Chemfiesta
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph Calculations Answers With Work Chemfiesta
 - Fact-Checking eBook Content of Ph Calculations Answers With Work Chemfiesta
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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