



Titration

Answer Key

Vocabulary: acid, analyte, base, dissociate, equivalence point, indicator, litmus paper, molarity, neutralize, pH, strong acid, strong base, titrant, titration, titration curve, weak acid, weak base

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

There are several definitions of acids and bases. According to the Brønsted-Lowry definition, an **acid** is a substance that is capable of donating a proton to another substance. A **base** is a substance that accepts protons. When an acid and a base are combined, the acid is **neutralized** as the base accepts the protons produced by the acid.



One way to determine if a solution is acidic or basic is to use **litmus paper**, as shown above. There are two types of litmus papers: red and blue.

How does litmus paper indicate an acid? *Both strips turn red.*

How does litmus paper indicate a neutral substance? *The red strip stays red and the blue strip stays blue.*

How does litmus paper indicate a base? *Both strips turn blue.*

Gizmo Warm-up

Litmus is an example of an **indicator**, a substance that changes color depending on its **pH** (pH is a measure of the concentration of protons, or H^+ ions). In the *Titration* Gizmo, you will use indicators to show how acids are neutralized by bases, and vice versa.

To begin, check that **1.00 M NaOH** is selected for the **Burette**, **Mystery HBr** is selected for the **Flask**, and **Bromthymol blue** is selected for the **Indicator**.



1. Look at the flask. What is the color of the bromthymol blue indicator? *Yellow*

2. What does this tell you about the pH of the solution in the flask? *pH is below 6.0*

Solutions with a pH below 7.0 are acidic, while those with a pH above 7.0 are basic.

3. Move the slider on the burette to the top to add about 25 mL of NaOH to the flask. What happens, and what does this tell you about the pH of the flask?

The solution turns blue, indicating the pH is now above 7.6.

Quick Titration Lab Answers

David Wild



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Robert Bruce Thompson, 2012-02-17. For students, DIY hobbyists, and science buffs who can no longer get real chemistry sets, this one of a kind guide explains how to set up and use a home chemistry lab with step-by-step instructions for conducting experiments in basic chemistry, not just to make pretty colors and stinky smells but to learn how to do real lab work. Purify alcohol by distillation. Produce hydrogen and oxygen gas by electrolysis. Smelt metallic copper from copper ore. You make yourself. Analyze the makeup of seawater, bone, and other common substances. Synthesize oil of wintergreen from aspirin and rayon fiber from paper. Perform forensics tests for fingerprints, blood, drugs, and poisons, and much more. From the 1930s through the 1970s, chemistry sets were among the most popular Christmas gifts, selling in the millions. But two decades ago, real chemistry sets began to disappear as manufacturers and retailers became concerned about liability. The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab, master laboratory skills, and work safely in your lab. The bulk of this book consists of 17 hands-on chapters that include multiple laboratory sessions on the following topics: Separating Mixtures, Solubility and Solutions, Colligative Properties of Solutions, Introduction to Chemical Reactions, Stoichiometry, Reduction/Oxidation, Redox Reactions, Acid-Base Chemistry, Chemical Kinetics, Chemical Equilibrium, and Le Chatelier's Principle, Gas Chemistry, Thermochemistry and Calorimetry, Electrochemistry, Photochemistry, Colloids and Suspensions, Qualitative Analysis, Quantitative Analysis, Synthesis of Useful

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