

# CHEMICAL EQUILIBRIUM



## EQUILIBRIUM CONSTANT 'K'.

For a General Reaction



The equilibrium constant expression is

$$K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$$

where  $K_c$  is the **Equilibrium Constant** (or  $K_p$  if they are all gases)

## RELATION BETWEEN $K_p$ AND $K_c$

For the Reaction



$$K_p = \frac{[P_c]^c \times [P_D]^d}{[P_A]^a \times [P_B]^b} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \frac{(RT)^{c+d}}{(RT)^{a+b}}$$

$$K_p = K_c (RT)^{\Delta n_g}$$

$$\text{if } \Delta n_g = 0 \implies K_p = K_c$$

Where,  $\Delta n_g = (c+d) - (a+b)$

= no. of moles of gaseous products - no. of moles of gaseous Reactants

## WHAT DOES THE VALUE OF 'K' MEAN ?

(a)  $K \gg 1$

Reactants



Products

If  $K \gg 1$ , the reaction is product-favoured; product predominates at Equilibrium.

(a)  $K \ll 1$

Reactants



Products

If  $K \ll 1$ , the reaction is reactant-favoured; reactant predominates at Equilibrium.

(a)  $K = 1$

The reaction lies in the middle (mix of reactants and products)

## MAGNITUDE OF 'K'

Small ( $K < 10^{-3}$ )



Reactants



Products

Mostly Reactants

Intermediate ( $10^{-3} \leq K \leq 10^3$ )



Reactants



Products

Significant amounts of reactants and products

Large ( $K > 10^3$ )



Reactants



Products

Mostly Products

# Phase Equilibrium Concept Review Answers Chapter 1

## section 4

**Y Pai**

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## **Table of Contents Phase Equilibrium Concept Review Answers Chapter 1section 4**

1. Understanding the eBook Phase Equilibrium Concept Review Answers Chapter 1section 4
  - The Rise of Digital Reading Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Advantages of eBooks Over Traditional Books
2. Identifying Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Phase Equilibrium Concept Review Answers Chapter 1section 4
  - User-Friendly Interface
4. Exploring eBook Recommendations from Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Personalized Recommendations
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 User Reviews and Ratings
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 and Bestseller Lists
5. Accessing Phase Equilibrium Concept Review Answers Chapter 1section 4 Free and Paid eBooks
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 Public Domain eBooks
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 eBook Subscription Services
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 Budget-Friendly Options

6. Navigating Phase Equilibrium Concept Review Answers Chapter 1section 4 eBook Formats
  - ePub, PDF, MOBI, and More
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 Compatibility with Devices
  - Phase Equilibrium Concept Review Answers Chapter 1section 4 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Highlighting and Note-Taking Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Interactive Elements Phase Equilibrium Concept Review Answers Chapter 1section 4
8. Staying Engaged with Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Phase Equilibrium Concept Review Answers Chapter 1section 4
9. Balancing eBooks and Physical Books Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Phase Equilibrium Concept Review Answers Chapter 1section 4
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Setting Reading Goals Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Fact-Checking eBook Content of Phase Equilibrium Concept Review Answers Chapter 1section 4
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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