

## Unit 6, Lesson 08: The pH of Salt Solutions, Answers

1. Complete the following chart:

Ion	Parent Acid or Base	Is the parent strong or weak?	Will this ion hydrolyze?	If the ion will hydrolyze (react with water), write the ionization reaction. If not, write "no reaction" or NR.	pH of ion in solution (acidic, basic or neutral)
$\text{NH}_4^{1+}$	$\text{NH}_3$	weak	yes	$\text{NH}_4^{1+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{NH}_3(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$	acidic
$\text{PO}_4^{3-}$	$\text{H}_3\text{PO}_3$	weak	yes	$\text{PO}_4^{3-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{HPO}_4^{2-}(\text{aq}) + \text{OH}^-(\text{aq})$	basic
$\text{Na}^{1+}$	$\text{NaOH}$	strong	no	NR	neutral
$\text{N}_2\text{H}_5^{1+}$	$\text{N}_2\text{H}_4$	weak	yes	$\text{N}_2\text{H}_5^{1+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{N}_2\text{H}_4(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$	acidic
$\text{HCO}_3^{1-}$	$\text{H}_2\text{CO}_3$	weak	yes	$\text{HCO}_3^{1-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{H}_2\text{CO}_3(\text{aq}) + \text{OH}^-(\text{aq})$	basic
$\text{SO}_4^{2-}$	$\text{H}_2\text{SO}_4$	strong	no	NR	neutral
$\text{Ca}^{2+}$	$\text{Ca}(\text{OH})_2$	strong	no	NR	neutral
$\text{NO}_3^{1-}$	$\text{HNO}_3$	strong	no	NR	neutral

### Notes and explanations:

- positive ions from weak bases are strong enough acids to hydrolyze. They will donate a proton to water and form the  $\text{H}_3\text{O}^+$  ion. Because this reaction forms the  $\text{H}_3\text{O}^+$  ion, it is a  $K_a$  reaction
- negative ions from weak acids are strong enough bases to hydrolyze. They will take a proton from water, leaving the  $\text{OH}^-$  ion behind. Because this reaction forms the  $\text{OH}^-$  ion, it is a  $K_b$  reaction

2. For each of the following salts:

- identify its parent acid and parent base and indicate if they are strong or weak

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**M Walker**



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## **Table of Contents Ph Of Salt Solutions Instructional Fair**

1. Understanding the eBook Ph Of Salt Solutions Instructional Fair
  - The Rise of Digital Reading Ph Of Salt Solutions Instructional Fair
  - Advantages of eBooks Over Traditional Books
2. Identifying Ph Of Salt Solutions Instructional Fair
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ph Of Salt Solutions Instructional Fair
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Of Salt Solutions Instructional Fair
  - Personalized Recommendations
  - Ph Of Salt Solutions Instructional Fair User Reviews and Ratings
  - Ph Of Salt Solutions Instructional Fair and Bestseller Lists
5. Accessing Ph Of Salt Solutions Instructional Fair Free and Paid eBooks
  - Ph Of Salt Solutions Instructional Fair Public Domain eBooks
  - Ph Of Salt Solutions Instructional Fair eBook Subscription Services
  - Ph Of Salt Solutions Instructional Fair Budget-Friendly Options

6. Navigating Ph Of Salt Solutions Instructional Fair eBook Formats
  - ePub, PDF, MOBI, and More
  - Ph Of Salt Solutions Instructional Fair Compatibility with Devices
  - Ph Of Salt Solutions Instructional Fair Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ph Of Salt Solutions Instructional Fair
  - Highlighting and Note-Taking Ph Of Salt Solutions Instructional Fair
  - Interactive Elements Ph Of Salt Solutions Instructional Fair
8. Staying Engaged with Ph Of Salt Solutions Instructional Fair
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ph Of Salt Solutions Instructional Fair
9. Balancing eBooks and Physical Books Ph Of Salt Solutions Instructional Fair
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ph Of Salt Solutions Instructional Fair
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ph Of Salt Solutions Instructional Fair
  - Setting Reading Goals Ph Of Salt Solutions Instructional Fair
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ph Of Salt Solutions Instructional Fair
  - Fact-Checking eBook Content of Ph Of Salt Solutions Instructional Fair
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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