

III. Determining the pH of various salt solutions.

During the experiment you would obtain a small amount of aqueous salts, and would measure their pH. You can predict whether a salt solution will be acidic, basic, or neutral following four situations:

1. A strong acid and a strong base will react to form a **neutral salt solution**.
2. A strong acid and a weak base yield a weakly **acidic salt solution**.
3. A weak acid and a strong base yield a weakly **basic salt solution**.
4. A weak acid plus a weak base can yield either **an acidic or basic solution**. This is the most complex of the four types of reactions. To predict whether a particular combination will be acidic or basic, tabulated K values of the conjugates must be compared. 4

Table 3. Determining the pH of various salt solutions

Salt	Name of salt	Salt type (acidic/basic)	Measured pH	Explanation for the difference between measured pH and 7.0
NaCl	Sodium chloride	Neutral salt	7.23	
<u>NaC₂H₃O₂</u>	Sodium acetate	Basic salt	9.41	
<u>NaHCO₃</u>	Sodium bicarbonate	Basic salt	9.29	
<u>Na₂CO₃</u>	Sodium carbonate	Basic salt	10.67	
<u>NH₄Cl</u>	Ammonium chloride	Acidic salt	5.79	
NH ₄ C ₂ H ₃ O ₂	Ammonium acetate	Basic salt	6.95	

How to write an explanation for the difference between measured pH and 7.0? You need to provide an equation for hydrolysis of each salt. This will confirm that there is a formation of OH⁻ or H₃O⁺

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North**



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