

of substance:

Substance:

of particles of a substance
in mol

6.02×10^{23} particles
amount of substance

Constant:

particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

Unit:

mol, measured in g
equivalent to M_r

Atomic Structure

Isotopes:

→ Same # of protons, same element
of p^+ and e^- but different
neutrons different M_r same

Physical Properties:

→ Can be different - e.g. different
melting + boiling points
→ i.e. heavy water has higher melting
point + boiling point + density

Chemical Properties:

→ Same no. of electrons + same electronic
structure in outer shell + transfer of e^- is
unaffected hence have same chemical
properties.
→ Different no. of n has no effect
on bonding + reactions

Compounds + ions FORMULAE:

Alkali M^+

Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure
- Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$

At RTP (20°C and 100kPa):
(293K and 100000Pa)

is $24\text{ dm}^3 \text{ mol}^{-1}$
at RTP.

Volume V
 $[\text{dm}^3]$

Electronegativity:

→ the measure of attraction of a bonded atom
in a covalent bond

ELECTRONEGATIVITY INCREASES ACROSS PERIOD:

- Increase in no. of protons in an atom
→ Nuclear charge increases

ELECTRONEGATIVITY DE:

- More electron shells
→ More shielding
→ Atomic radius increases
→ Nuclear charge decreases
→ Nuclear attraction decreases
→ Electronegativity decreases

OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS



→ Consists of a small dense
nucleus of protons + neutrons
surrounded by electrons orbiting
in electron shells in shells

- No. of p^+ = No. of e^-

→ Equal + opposite charges cancel
in atom has neutral charge
- Neutrons important in nuclear
stability - Electrostatic repulsion
between p^+ in small dense nucleus

Mass Number:

→ No. of p^+ + n in nucleus of atom
- p^+ and n have relative mass of 1
and e^- mass is negligible is 0

Relative Isotopic Mass:

→ The mass of an atom of an isotope
relative to $1/12$ th the mass of an atom of $C-12$

Relative Atomic Mass:

→ The weighted mean mass of an atom of
an element relative to $1/12$ th the mass of
an atom of $C-12$

Weighted Mean Mass:

→ Average mass of all isotopes, taking into
account the relative abundance of each isotope

- It dipoles
Molecule is
out + inlets

e.g. H_2O

- O is more electronegative than H
- O-H has permanent dipole since do not cancel there is an over

e.g. CO_2

O is more electronegative than C
 $\text{C}=\text{O}$ has a dipole
Molecule bond

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Julia Schneider



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1. Understanding the eBook Ocr Chemistry A Level Unit 2 2014 Paper
 - The Rise of Digital Reading Ocr Chemistry A Level Unit 2 2014 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A Level Unit 2 2014 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry A Level Unit 2 2014 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A Level Unit 2 2014 Paper
 - Personalized Recommendations
 - Ocr Chemistry A Level Unit 2 2014 Paper User Reviews and Ratings
 - Ocr Chemistry A Level Unit 2 2014 Paper and Bestseller Lists
5. Accessing Ocr Chemistry A Level Unit 2 2014 Paper Free and Paid eBooks
 - Ocr Chemistry A Level Unit 2 2014 Paper Public Domain eBooks
 - Ocr Chemistry A Level Unit 2 2014 Paper eBook Subscription Services
 - Ocr Chemistry A Level Unit 2 2014 Paper Budget-Friendly Options
6. Navigating Ocr Chemistry A Level Unit 2 2014 Paper eBook Formats

- ePub, PDF, MOBI, and More
 - Ocr Chemistry A Level Unit 2 2014 Paper Compatibility with Devices
 - Ocr Chemistry A Level Unit 2 2014 Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Ocr Chemistry A Level Unit 2 2014 Paper
 - Highlighting and Note-Taking Ocr Chemistry A Level Unit 2 2014 Paper
 - Interactive Elements Ocr Chemistry A Level Unit 2 2014 Paper
8. Staying Engaged with Ocr Chemistry A Level Unit 2 2014 Paper
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry A Level Unit 2 2014 Paper
9. Balancing eBooks and Physical Books Ocr Chemistry A Level Unit 2 2014 Paper
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry A Level Unit 2 2014 Paper
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry A Level Unit 2 2014 Paper
- Setting Reading Goals Ocr Chemistry A Level Unit 2 2014 Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry A Level Unit 2 2014 Paper
- Fact-Checking eBook Content of Ocr Chemistry A Level Unit 2 2014 Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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