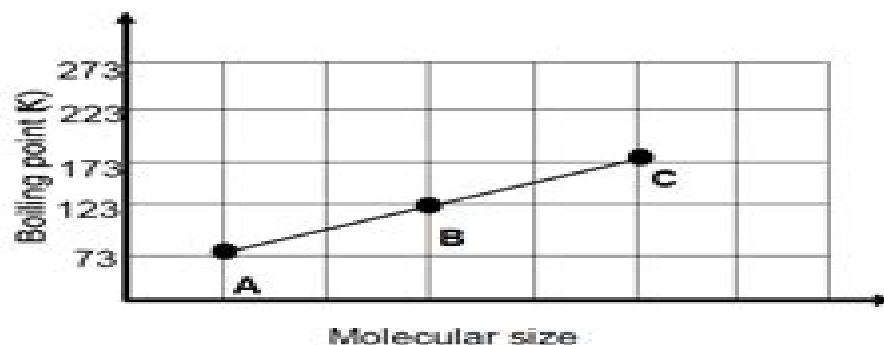


QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



2.1.1 Define the term *boiling point*. (2)

2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)

2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)

2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

2.2.1 Determine the molecular mass (x) of phosphine (1)

2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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Table of Contents Physical Science Grade Caps Question Papers

1. Understanding the eBook Physical Science Grade Caps Question Papers
 - The Rise of Digital Reading Physical Science Grade Caps Question Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade Caps Question Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade Caps Question Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade Caps Question Papers
 - Personalized Recommendations

- Physical Science Grade Caps Question Papers User Reviews and Ratings
- Physical Science Grade Caps Question Papers and Bestseller Lists
- 5. Accessing Physical Science Grade Caps Question Papers Free and Paid eBooks
 - Physical Science Grade Caps Question Papers Public Domain eBooks
 - Physical Science Grade Caps Question Papers eBook Subscription Services
 - Physical Science Grade Caps Question Papers Budget-Friendly Options
- 6. Navigating Physical Science Grade Caps Question Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Grade Caps Question Papers Compatibility with Devices
 - Physical Science Grade Caps Question Papers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Grade Caps Question Papers
 - Highlighting and Note-Taking Physical Science Grade Caps Question Papers
 - Interactive Elements Physical Science Grade Caps Question Papers
- 8. Staying Engaged with Physical Science Grade Caps Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade Caps Question Papers
- 9. Balancing eBooks and Physical Books Physical Science Grade Caps Question Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade Caps Question Papers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Grade Caps Question Papers
 - Setting Reading Goals Physical Science Grade Caps Question Papers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Grade Caps Question Papers
 - Fact-Checking eBook Content of Physical Science Grade Caps Question Papers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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