Molar Concentration

Part A: Describing Molar Concentration

Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?

1.5 mol of Nacl if you need a certain countries

2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate. KNO2 : - - = 101,119/mel 0,45 mel KNOL

V= 01225L

0,45 mm 1003 = 2 M

Find the molarity of a solution made from 275 g of CuSO₄ dissolved in enough

water to make 4.25 L. 2755 = 1.72 wil 6.504

275y 6.507

154.45 = 1.72 wil 6.504

1355 | 154.655/mil | 1.72 wil = 0.40 M 6.504

4.25 L

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

5.159 Te = 0.02 mol IE

0,02 - IZ 0.089 M IZ

5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

was one Watch ! 58,449

22.49 58,449 = 0.039 mol Macl

0.089 mol = 0.36 L

Part A Describing Molar Concentration

B Lingard

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