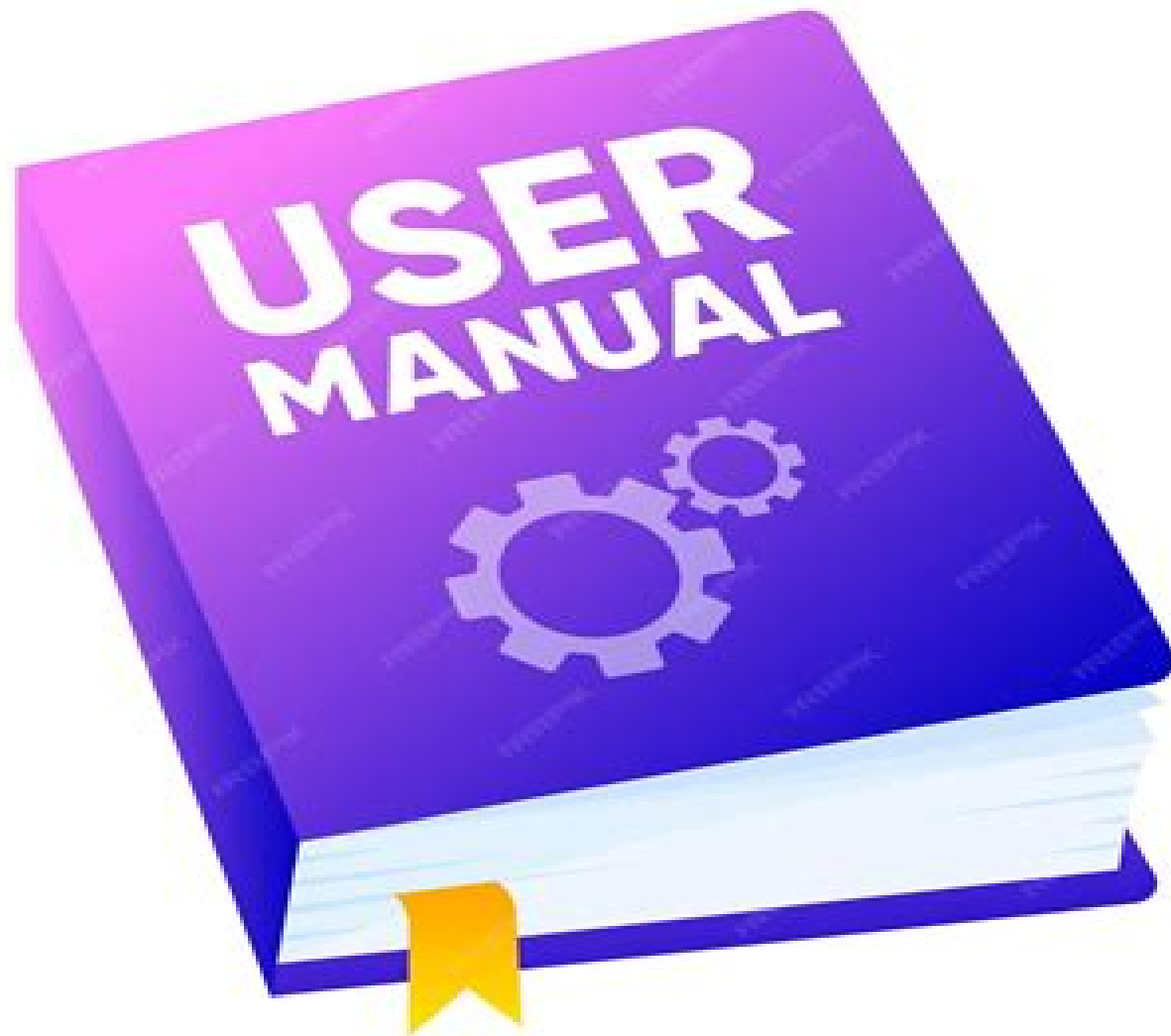




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reaction a redox reaction explain your answer $2\text{NaCl aq} + \text{Pb} \rightarrow \text{Pb}^{2+} + 2\text{Cl}^{-}$ $2\text{NaNO}_3 \rightarrow 2\text{NaNO}_2 + \text{O}_2$ $\text{PbCl}_2 \rightarrow \text{Pb} + \text{Cl}_2$

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 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 $5\text{Co} + \text{I}_2 \rightarrow 5\text{CoI}_2$ $\text{I}_2 \rightarrow 2\text{I}^{+5}$ $5\text{Co} \rightarrow 5\text{Co}^{+2}$ $5\text{Co} \rightarrow 5\text{Co}^{+2}$

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