

# CHEMISTRY UNIT 2 REVISION

## Energetics:

→ How to calculate general bond enthalpy?

- Bonds broken – bonds formed = H

→ What are the products of combustion?

- Carbon dioxide and water

→ Enthalpy change of formation conditions?

- One mole
- Standard states

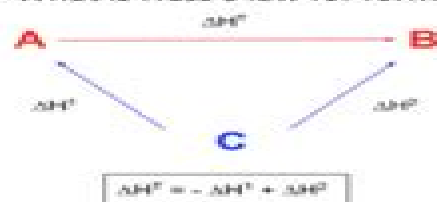
→ Enthalpy change of atomisation conditions?

- One mole
- Standard conditions
- One mole of gaseous molecule forms (product)

→ What is Hess's law for combustion?



→ What is Hess's law for formation?



→ How do you use Hess's law to solve equations?

- Check whether its formation or combustion and draw accordingly
- Write the enthalpies for each product or reactant according to the arrows
- Keep in mind A is the elements or compounds used, B is the desired product, C is waste product

→ How to calculate enthalpy change?

- $Q = mct$ , so  $H = Q / n$

→ How to increase thermal stability?

- Decrease ionic charge
- Increase ionic radius

# October November 2chemistry Paper 60620

**Judd E. Hollander**



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