

**Chapter 12: Atoms and the Periodic Table**  
**Subatomic Particles**

Three fundamental particles of the atom are the \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_. At the center of each atom lies the atomic \_\_\_\_\_, which consists of \_\_\_\_\_ and \_\_\_\_\_. The atomic number refers to the number of \_\_\_\_\_ in the nucleus. All atoms of the same element have the same number of \_\_\_\_\_, hence, the same atomic number.

Isotopes are atoms that have the same number of \_\_\_\_\_ but a different number of \_\_\_\_\_. An isotope is identified by its atomic mass number, which is the total number of \_\_\_\_\_ and \_\_\_\_\_ in the nucleus. A carbon isotope that has 6 \_\_\_\_\_ and 6 \_\_\_\_\_ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 \_\_\_\_\_ and 8 \_\_\_\_\_, on the other hand, is carbon-14.

1. Complete the following table:

Isotope	Number of...		
	Electrons	Protons	Neutrons
Hydrogen-1	1		
Chlorine-36		17	
Nitrogen-14			7
Potassium-40	19		
Arsenic-75		33	
Gold-197			118

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



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## **Table of Contents Physical Science Chapter 12 Answers**

1. Understanding the eBook Physical Science Chapter 12 Answers
  - The Rise of Digital Reading Physical Science Chapter 12 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 12 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Physical Science Chapter 12 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 12 Answers
  - Personalized Recommendations

- Physical Science Chapter 12 Answers User Reviews and Ratings
- Physical Science Chapter 12 Answers and Bestseller Lists
- 5. Accessing Physical Science Chapter 12 Answers Free and Paid eBooks
  - Physical Science Chapter 12 Answers Public Domain eBooks
  - Physical Science Chapter 12 Answers eBook Subscription Services
  - Physical Science Chapter 12 Answers Budget-Friendly Options
- 6. Navigating Physical Science Chapter 12 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science Chapter 12 Answers Compatibility with Devices
  - Physical Science Chapter 12 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Chapter 12 Answers
  - Highlighting and Note-Taking Physical Science Chapter 12 Answers
  - Interactive Elements Physical Science Chapter 12 Answers
- 8. Staying Engaged with Physical Science Chapter 12 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Chapter 12 Answers
- 9. Balancing eBooks and Physical Books Physical Science Chapter 12 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Chapter 12 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chapter 12 Answers
  - Setting Reading Goals Physical Science Chapter 12 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chapter 12 Answers
  - Fact-Checking eBook Content of Physical Science Chapter 12 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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