

# PHYSICAL CHEMISTRY

## (Complete Syllabus)

### SECTION - A

#### Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

---

Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces  $x$  gm of  $\text{Fe}_2\text{O}_3$ . Suppose molecular weight of  $\text{O}_2 = y$ , molecular weight of  $\text{Fe}_2\text{O}_3 = z$   
(1)  $\frac{2xz}{3y}$  (2)  $\frac{3xy}{2z}$   
(3)  $\frac{2y}{3xz}$  (4)  $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be  
(1)  $0.02 N_0$  (2)  $0.16 N_0$   
(3)  $0.32 N_0$  (4)  $0.08 N_0$
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with  $\text{Na}_2\text{CO}_3$  to precipitate the calcium as calcium carbonate. This  $\text{CaCO}_3$  is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of  $\text{CaCl}_2$  in the original mixture is  
(1) 15.2%  
(2) 32.1%  
(3) 21.8%  
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of  $\text{I}_2$  liberated is  
(1) 25.4 g (2) 12.7 g  
(3) 2.54 g (4) 1.27 g
- $x_1$  ml solution having normality 1 is diluted to get a solution of normality  $y_2$ . What would be volume required for dilution?  
(1)  $x_1 \left(1 - \frac{y_1}{y_2}\right)$  (2)  $x_1 \left(\frac{y_2}{y_1} - 1\right)$   
(3)  $x_1 \left(\frac{y_1}{y_2} - 1\right)$  (4)  $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound  $\text{ICl}_x$  was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is  
(1)  $\text{ICl}$  (2)  $\text{ICl}_3$   
(3)  $\text{ICl}_5$  (4)  $\text{ICl}_7$
- If 12.25 g of a sample of  $\text{KClO}_3$  of 75% purity and rest vapourisable impurities is strongly heated to give  $\text{KCl}$ ,  $\text{KClO}_4$  and 0.32 gm  $\text{O}_2$ . The mass of residue obtained is  
(1) 1.769 g (2) 7.096 g  
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is  
(1)  $\text{XeF}_2$  (2)  $\text{XeF}_4$   
(3)  $\text{XeF}_6$  (4)  $\text{XeF}_8$
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

# Physical Chemistry Guide University Questions And Answers

**Daniel F McAuley**



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## **Table of Contents Physical Chemistry Guide University Questions And Answers**

1. Understanding the eBook Physical Chemistry Guide University Questions And Answers
  - The Rise of Digital Reading Physical Chemistry Guide University Questions And Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry Guide University Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Chemistry Guide University Questions And Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry Guide University Questions And Answers
  - Personalized Recommendations
  - Physical Chemistry Guide University Questions And Answers User Reviews and Ratings
  - Physical Chemistry Guide University Questions And Answers and Bestseller Lists
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  - Interactive Elements Physical Chemistry Guide University Questions And Answers
8. Staying Engaged with Physical Chemistry Guide University Questions And Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Chemistry Guide University Questions And Answers
9. Balancing eBooks and Physical Books Physical Chemistry Guide University Questions And Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Chemistry Guide University Questions And Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
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- Setting Reading Goals Physical Chemistry Guide University Questions And Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Chemistry Guide University Questions And Answers
- Fact-Checking eBook Content of Physical Chemistry Guide University Questions And Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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