

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion always forms with a charge of 2-? _____ (anion)
which cation is always with a charge of 2+? _____ (cation)
2. How electrons in the halogen group has _____ valence electrons, and needs to _____ 1 more to have 8. This is because of the _____ rule.
3. How _____ bond in the alkaline elements oppositely charged ions.
How is ionic bond is _____
4. On a molecular level, atoms _____ molecules, formed by an ionic bond, atoms _____ molecules.
5. Elements group 16 has 6 valence electrons. It is usually found as _____, and atoms group 17 usually found as _____.
How can group 17 be distinguished, in terms, other elements, or molecules formed by ionic bonding? _____
6. How atoms can have up to _____ valence electrons. This usually happens in the periodic table that all have 8 valence electrons, is called _____
7. Atoms can gain, lose, or share _____ by forming ionic bonds.

B. Formulas for Ionic Compounds - Fill in the chart below

Write the correct chemical formula for	Magnesium	Fluorine	Phosphorus	Chlorine
1. Lithium				
2. Potassium				
3. Sodium				

C. Write the chemical formulas for each of the following ions:

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if they combine numbers already given.

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D Kirk



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