

Chemistry

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(Chapter - 3) (Chemical Kinetics) (Exercises)
(Class 12)

Question 3.1:

From the rate expression for the following reactions, determine their order of reaction and the dimensions of the rate constants.



Answer 3.1:



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{NO}]^2} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^2} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^2\text{L}^{-2}} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of the reaction = 2

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{H}_2\text{O}_2][\text{I}^-]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})(\text{mol L}^{-1})} = \text{L mol}^{-1}\text{s}^{-1}$$



Therefore, order of reaction = $3/2$

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{CH}_3\text{CHO}]^{3/2}} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{(\text{mol L}^{-1})^{3/2}} = \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol}^{3/2}\text{L}^{3/2}} = \text{L}^{3/2}\text{mol}^{1/2}\text{s}^{-1}$$



Therefore, order of the reaction = 1

$$\text{Dimension of } k = \frac{\text{Rate}}{[\text{C}_2\text{H}_5\text{Cl}]} \Rightarrow \frac{\text{mol L}^{-1}\text{s}^{-1}}{\text{mol L}^{-1}} = \text{s}^{-1}$$

Question 3.2:

For the reaction:



The rate = $k[\text{A}][\text{B}]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of reaction when $[\text{A}] = 0.1 \text{ mol L}^{-1}$, $[\text{B}] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[\text{A}]$ is reduced to 0.06 mol L^{-1} .

Answer 3.2:

The initial rate of reaction is

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8.0 \times 10^{-9} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$$

When $[\text{A}]$ is reduced from 0.1 mol L^{-1} to 0.06 mol L^{-1} ,

$$\text{The concentration of A reacted} = (0.1 - 0.06) \text{ mol L}^{-1} = 0.04 \text{ mol L}^{-1}$$

$$\text{Therefore, concentration of B reacted} = \frac{1}{2} \times 0.04 \text{ mol L}^{-1}$$

$$\text{Then concentration of B available, } [\text{B}] = (0.2 - 0.02) \text{ mol L}^{-1} = 0.18 \text{ mol L}^{-1}$$

After $[\text{A}]$ is reduced to 0.06 mol L^{-1} , the rate of reaction is given by,

$$\text{Rate} = k[\text{A}][\text{B}]^2 \Rightarrow (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1})(0.06 \text{ mol L}^{-1})(0.18 \text{ mol L}^{-1})^2 = 3.89 \text{ mol L}^{-1} \text{ s}^{-1}$$

Question 3.3:

The decomposition of NH_3 on platinum is zero order reaction. What are the rates of production of N_2 and H_2 if $k = 2.5 \times 10^{-4} \text{ mol}^{-1} \text{ L s}^{-1}$?

Answer 3.3:

The decomposition of NH_3 on platinum surface is represented by the following equation.



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1. Understanding the eBook Part2solution Chapters 1cccd
 - The Rise of Digital Reading Part2solution Chapters 1cccd
 - Advantages of eBooks Over Traditional Books
2. Identifying Part2solution Chapters 1cccd
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Part2solution Chapters 1cccd
 - User-Friendly Interface
4. Exploring eBook Recommendations from Part2solution Chapters 1cccd
 - Personalized Recommendations

- Part2solution Chapters 1cccd User Reviews and Ratings
- Part2solution Chapters 1cccd and Bestseller Lists
- 5. Accessing Part2solution Chapters 1cccd Free and Paid eBooks
 - Part2solution Chapters 1cccd Public Domain eBooks
 - Part2solution Chapters 1cccd eBook Subscription Services
 - Part2solution Chapters 1cccd Budget-Friendly Options
- 6. Navigating Part2solution Chapters 1cccd eBook Formats
 - ePub, PDF, MOBI, and More
 - Part2solution Chapters 1cccd Compatibility with Devices
 - Part2solution Chapters 1cccd Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Part2solution Chapters 1cccd
 - Highlighting and Note-Taking Part2solution Chapters 1cccd
 - Interactive Elements Part2solution Chapters 1cccd
- 8. Staying Engaged with Part2solution Chapters 1cccd
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Part2solution Chapters 1cccd
- 9. Balancing eBooks and Physical Books Part2solution Chapters 1cccd
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Part2solution Chapters 1cccd
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Part2solution Chapters 1cccd
 - Setting Reading Goals Part2solution Chapters 1cccd
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Part2solution Chapters 1cccd
 - Fact-Checking eBook Content of Part2solution Chapters 1cccd

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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