

Polyatomic Ions Worksheet

Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

| | | | |
|--------------|-------|---------------------|-------|
| sulfate | _____ | CO_3^{2-} | _____ |
| nitrite | _____ | MnO_4^{1-} | _____ |
| perphosphate | _____ | SO_3^{2-} | _____ |
| hypiodite | _____ | BrO_2^{1-} | _____ |
| chlorite | _____ | CO_3^{2-} | _____ |
| phosphite | _____ | PO_3^{2-} | _____ |
| percarbonate | _____ | ClO^{1-} | _____ |
| bromate | _____ | IO_3^{1-} | _____ |
| hyposulfite | _____ | PO_4^{2-} | _____ |
| permanganate | _____ | NO_3^{1-} | _____ |
| carbonate | _____ | SO_4^{2-} | _____ |

2. Name or write the formula for the following Type I polyatomic ionic compounds

| | | | |
|-------------------------|-------|---|-------|
| beryllium hydroxide | _____ | $\text{Ba}(\text{IO}_3)_2$ | _____ |
| sodium nitrite | _____ | $\text{Ga}(\text{CNO})_3$ | _____ |
| ammonium chloride | _____ | Ag_2SO_4 | _____ |
| calcium bisulfate | _____ | MgCO | _____ |
| rubidium perchlorate | _____ | NH_2NO_2 | _____ |
| strontium sulfite | _____ | $\text{Al}(\text{C}_2\text{H}_3\text{O}_5)_3$ | _____ |
| aluminum acetate | _____ | SrSO_4 | _____ |
| ammonium nitrate | _____ | RbClO_2 | _____ |
| magnesium hypocarbonite | _____ | $\text{Ca}(\text{HSO}_4)_2$ | _____ |
| silver hyposulfite | _____ | NH_4Cl | _____ |
| gallium cyanate | _____ | NaNO | _____ |
| barium iodate | _____ | $\text{Be}(\text{OH})_2$ | _____ |

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M W Roberts, John M Thomas



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and some evening sessions where the participants were offered the opportunity to present their own work and discuss their ideas with senior scientists. Since the Institute took place a few weeks after Prof Dr John C Slater was suddenly taken from our scientific community it was a great honor for us to dedicate this course on behalf of the organizing committee to the late John C Slater.

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