

Name _____ Date _____

PHYSICAL & CHEMICAL CHANGE

In a physical change, the original substance still exists; it has only changed in form (i.e. size, shape, state of matter, etc). In addition, a physical change is always reversible - you can get back what you started with! In a chemical change, however, a completely new substance is produced - you can never get back what you started with! Energy (heat) changes typically accompany chemical changes, but can sometimes occur in physical changes as well. Do not assume a chemical change is taking place just because energy is involved! Think: Do I have a totally different substance than I started with? If so, this is always a chemical change.

Read the following statements. Classify the following as being a physical or chemical change.

PHYSICAL CHANGE	CHEMICAL CHANGE	STATEMENTS
		An ice cube melts on a hot summer day.
		Green grass making its own food.
		Hydrochloric acid reacts with potassium hydroxide to produce salt, water, and heat.
		Pancakes cook on a griddle.
		non-metals do not conduct heat or electricity well
		On a cold fall morning, water vapor in the atmosphere condenses into dew on the grass.
		A tire is deflated (the air is let out).
		Potassium chlorate decomposes to potassium chloride and oxygen gas.
		When placed in water, a sodium pellet catches on fire as hydrogen gas is given off and sodium hydroxide forms.
		Wood rots out in the rain.
		An iron nail rusts when left out in the rain.
		A pellet of sodium is sliced in two.
		Food is digested in the stomach
		Milk sours after being kept past the expiration date.
		A pellet of sodium is sliced in two.
		Sugar dissolves in water.
		Sodium hydroxide dissolves in water.

Physical Change Vs Chemical Change Worksheet

Chemistry If8766

Zhe Xu



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Physical and Chemical Changes (eBook) Edward P. Ortleb, Richard Cadice, 1993-09-01 This book presents a program of basic studies in physical and chemical changes of matter. The definition of matter is presented along with explanations of states and properties of matter. Topics include atoms, molecules, elements, compounds, mixtures, solutions, symbols, and formulas. Each of the twelve teaching units in this book is introduced by a color transparency, print books, or PowerPoint slide eBooks that emphasizes the basic concept of the unit and presents questions for discussion. Reproducible student pages provide reinforcement and follow-up activities. The teaching guide offers descriptions of the basic concepts to be presented, background information, suggestions for enrichment activities, and a complete answer key. [Properties of Matter: Physical Changes vs. Chemical Changes Gr. 5-8](#) George Graybill, 2015-09-01 This is the chapter slice Physical Changes vs Chemical Changes from the full lesson plan Properties of Matter. Discover what matter is and is not. Learn about and the difference between a mixture and a solution. Chocked full with hands-on activities to understand the various physical and chemical changes to matter. Our resource provides ready-to-use information and activities for remedial students using simplified language and vocabulary. Written to grade, these science concepts are presented in a way that makes them more accessible to students and easier to understand. Our resource is jam-packed with experiments, reading passages, and activities all for students in grades 5 to 8. Color mini-posters and answer key included and can be used effectively for test prep and your whole class. All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and STEM initiatives.

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During a chemical change the chemical properties of a substance change and one or more new substances form Every chemical reaction involves a transfer of energy During a chemical reaction atoms in matter are rearranged but no atoms are gained or lost

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Physical and Chemical Change Alan V. Jones, Nigel Heslop, Mike Clemmet, 1995

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