

CBSE Class 10 Science Sample Paper SA 1

Set 4

Time: 3hrs

M.M.: 80

General Instructions :-

- (i) The question paper comprises of two Sections, A and B. You are to attempt both the sections.
- (ii) All questions are compulsory.
- (iii) There is no overall choice. However, internal choice has been provided in all the five questions of five marks category. Only one option in such questions is to be attempted.
- (iv) All questions of Section-A and all questions of Section-B are to be attempted separately.
- (v) Question numbers 1 to 3 in Section-A are one mark questions. These are to be answered in one word or in one sentence.
- (vi) Question numbers 4 to 7 in Section-A are two marks questions. These are to be answered in about 30 words each.
- (vii) Question numbers 8 to 19 in Section-A are three marks questions. These are to be answered in about 50 words each.
- (viii) Question numbers 20 to 24 in Section-A are five marks questions. These are to be answered in about 70 words each.
- (ix) Question numbers 25 to 42 in Section-B are multiple choice questions based on practical skills. Each question is a one mark question. You are to select one most appropriate response out of the four provided to you.

SECTION: 'A'

1. In domestic electric circuit, with which wire do we connect a fuse? [1]
2. Name the two components of central nervous system in humans. [1]
3. Name one fuel used in nuclear reactor. [1]
4. Write the chemical reaction for respiration. Why is it considered an exothermic reaction? [2]
5. Mention one physical property each of a metal and a non-metal which is exceptional to their general properties. [2]
6. A battery of 12 V is connected to a series combination of resistors, 3 Ω , 4 Ω , 5 Ω , and 12 Ω . How much current would flow through the 12 Ω resistor? [2]
7. Why do two magnetic field lines never intersect each other? [2]
8. In the electrolysis of water: [3]
 - (i) Name the gas collected at the cathode and anode respectively
 - (ii) Why is the volume of one gas collected at one electrode double than that at the other? Name this gas.
 - (iii) How will you test the evolved gases?

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds Cl^- aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{Ag}^+ \text{aq} + 2 \text{F}^- \text{aq} + 2 \text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2 \text{NH}_4^+ \text{aq} + 2 \text{F}^- \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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