

Annotated periodic table of the elements

Group names*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Annotated periodic table of the elements

Group names*

1 Hydrogen & the alkali metals	14 Crystallogens
2 Alkaline earth metals	15 Pnictogens
11 Coinage metals	16 Chalcogens
12 Volatile metals	17 Halogens
13 Ionogens	18 Helium group

*Groups 1 to 10 are named after their

first members i.e. Group 1 is the Lanthanum Group

Most common oxidation states, electronegativity (after Pauling)

Atomic number and symbol
Atomic weight and name

57 to 70

89 to 102

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H 1.01 Hydrogen												5 B 10.81 Boron	6 C 12.01 Carbon	7 N 14.01 Nitrogen	8 O 16.00 Oxygen	9 F 18.99 Fluorine	10 Ne 20.18 Neon
3 Li 6.94 Lithium	4 Be 9.012 Beryllium											11 Na 22.99 Sodium	12 Mg 24.31 Magnesium				
19 K 39.10 Potassium	20 Ca 40.08 Calcium											27 Co 58.93 Cobalt	28 Ni 58.69 Nickel	29 Cu 63.55 Copper	30 Zn 65.38 Zinc	31 Ga 69.72 Gallium	32 Ge 72.64 Germanium
37 Rb 85.47 Rubidium	38 Sr 87.62 Strontium											47 Ag 107.87 Silver	48 Cd 112.41 Cadmium	49 In 114.82 Indium	50 Sn 118.71 Tin	51 Sb 121.76 Antimony	52 Te 127.60 Tellurium
55 Cs 132.91 Cesium	56 Ba 137.33 Barium											67 Lu 174.96 Lutetium	68 Yb 173.05 Ytterbium	69 Tl 204.38 Thallium	70 Pb 207.2 Lead	71 Bi 208.98 Bismuth	72 Po 209 Polonium
87 Fr 223.02 Francium	88 Ra 226.025 Radium											103 Nh 289.10 Nihonium	104 Fl 289.10 Flerovium	105 Mc 289.10 Moscovium	106 Lv 289.10 Livermorium	107 Ts 289.10 Tennessine	108 Og 289.10 Oganesson

Most common oxidation states, electronegativity (Allen-Pauling)

Atomic number and symbol, atomic weight and name

CLASSIFICATION

Metals	Nonmetals
Pre-transition	
Active	Corrosive
Alkaline earth	
Alkaloids	
Transition	Intermediate
Post-transition	
Metals	
Metals (Ru-Pt, Os-Au)	Metals gas

57-70 La 138.91 Lanthanum	71-80 Ce 140.12 Cerium	81-90 Pr 140.91 Praseodymium	91-100 Nd 144.24 Neodymium	101-110 Pm 144.91 Promethium	111-120 Sm 150.36 Samarium	121-130 Eu 151.96 Europium	131-140 Gd 157.25 Gadolinium	141-150 Tb 158.93 Terbium	151-160 Dy 162.50 Dysprosium	161-170 Ho 164.93 Holmium	171-180 Er 167.26 Erbium	181-190 Tm 168.93 Thulium	191-200 Yb 173.05 Ytterbium
89-102 Ac 227.03 Actinium	103-112 Th 232.04 Thorium	113-122 Pa 231.04 Protactinium	123-132 U 238.03 Uranium	133-142 Np 237.05 Neptunium	143-152 Pu 244.06 Plutonium	153-162 Am 243.06 Americium	163-172 Cm 247.07 Curium	173-182 Bk 247.07 Berkelium	183-192 Cf 251.08 Californium	193-202 Es 252.08 Einsteinium	203-212 Fm 257.10 Fermium	213-222 Md 258.10 Mendelevium	223-232 No 259.10 Nobelium

Radioactive

Approximate isotope is for either a metalloid or a post-transition metal

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalyst properties. Very weakly electropositive, inert to acids.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and tellurates.

Metals

Nonmetals

Metalloids

- Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colourless gases, C, P, and Se are metalloids looking solids while Si is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- *as black P, the most stable form in ambient conditions
- Group 1**
This starts as Li-Li, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lu. Trends going down Lu-Lu, and the La branch, resemble those of groups 1-2.
- *Lu branch, more resemble those of groups 1-2

Corrosive nonmetals

- O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colourless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answers

**Denise Guinn, Rebecca Brewer, Rachel
C. Lum**

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

GCE O Level Examination Past Papers with Answer Guides: Chemistry India Edition Cambridge International Examinations, 2003-12-16 These Collections Of The Official Past Papers Of The Gce O Level Examinations From The University Of Cambridge International Examinations Has Been Developed For Students Of Gce O Level These Books Will Act As Tools For Preparation And Revision For Students These Books Have An Edited Answer Guide For Each Paper Based On The Marks Scheme Written By Cie Principal

Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011

Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts After a brief review of each section s critical ideas students are taken through stepped out worked examples try it yourself examples and chapter quizzes all structured to reinforce chapter objectives and build problem solving techniques The solutions manual includes detailed solutions to all odd numbered exercises in the text

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