

Periodic Trends

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

5. Order the following elements by increasing atomic mass:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing atomic number:

Li, Na, K, Rb, Cs, Fr

7. Order the following elements by increasing atomic weight:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing atomic density:

Li, Na, K, Rb, Cs, Fr

9. Order the following elements by increasing atomic volume:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing atomic mass number:

Li, Na, K, Rb, Cs, Fr

Name: _____ Date: _____ Period: _____

Periodic Trends Test Study Guide

Section I: Review and Recall of the Periodic Table

1. What are the groups (families) of elements in the periodic table?

- 1. s-block elements (groups 1 and 2)
- 2. p-block elements (groups 13-18)
- 3. d-block elements (transition metals, groups 3-10)
- 4. f-block elements (lanthanides and actinides, groups 14-16)

2. Label the following families on the periodic table: s-block, p-block, d-block, f-block, transition metals, main group elements, noble gases, halogens, alkali metals, alkaline earth metals, lanthanides, actinides.

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Periodic Table And Chemistry Study Guide Answers

Roman Wölfel



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