

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Molar masses, stoichiometric ratios, limiting reagent, percent yield, theoretical yield, actual yield, percent error.

Chemists will often make reactions in which one reactant is in excess and the other is limiting. The limiting reagent is the one that runs out first.



The ratio required for the components for the reaction is:

$\text{H}_2(\text{g})$	200.00 grams
$\text{O}_2(\text{g})$	50.00 grams
$\text{H}_2\text{O}(\text{l})$	180.00 grams
H_2O	180.00 grams

1. 200.00 g of $\text{H}_2(\text{g})$ and 50.00 g of $\text{O}_2(\text{g})$ are used. How much $\text{H}_2\text{O}(\text{l})$ is produced?
 How much $\text{H}_2(\text{g})$ is left over? How much $\text{O}_2(\text{g})$ is left over?



2. How much $\text{H}_2\text{O}(\text{l})$ is produced? $\text{H}_2(\text{g})$ is the limiting reagent.

3. How much $\text{H}_2(\text{g})$ is left over? $\text{O}_2(\text{g})$ is the limiting reagent.



4. How much $\text{H}_2\text{O}(\text{l})$ is produced? $\text{O}_2(\text{g})$ is the limiting reagent.



Percent Yield And Limiting Reactants Quiz

Kathleen Armour



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