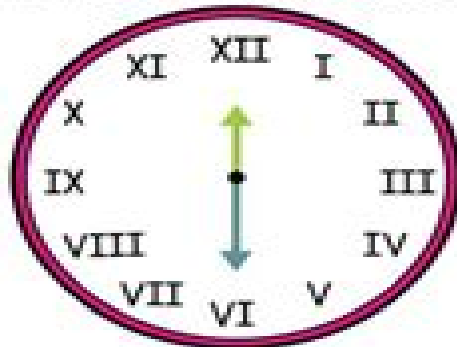


# ROMAN NUMERAL CHARTS

## ROMAN NUMERALS



I = 1	VII = 7
II = 2	VIII = 8
III = 3	IX = 9
IV = 4	X = 10
V = 5	XI = 11
VI = 6	XII = 12



## ROMAN NUMERALS

SYMBOLS

From 1 - 100

1	5	10	50	100
I	V	X	L	C

COMBINATIONS From 10 to 100

10	20	30	40	50
X	XX	XXX	XL	L
60	70	80	90	100
LX	LXX	LXXX	XC	C

HOW TO CONVERT

25

XXV

75

LXXV

150

CL

A letter placed AFTER a bigger value adds to it.

VI = 5 + 1 = 6

A letter placed BEFORE a bigger value subtracts from it.

IV = 5 - 1 = 4

# Roman Numeral Study Guide

**J Spring**



## **Roman Numeral Study Guide:**

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 $2 \text{ CO}_2(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 2 \text{ O}_2(\text{g}) \quad \Delta H = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 3 \text{ H}_2(\text{g}) \quad \Delta H = 283 \text{ kJ}$

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web chemistry 120 hess s law worksheet calculate  $\Delta H$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + 2 \text{ C}_2\text{H}_6(\text{g}) \rightarrow 3 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta H = 1411 \text{ kJ}$  mole  $\text{C}_2\text{H}_6(\text{g})$   $2 \text{ O}_2(\text{g}) + 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta H = 1560 \text{ kJ}$  mole  $\text{H}_2(\text{g})$   $2 \text{ O}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$   $\Delta H = 285 \text{ kJ}$  mole

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combustion of magnesium oxide  $\text{Mg}$  s 1  $2\text{O}_2$  g  $\text{MgO}$  s we used

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows 1  $\text{NaOH}$  s  $\text{Na}$  aq  $\text{OH}^-$  aq

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web hess s law states that if a reaction is the sum of two or more other reactions the  $\Delta H$  for the overall process must be the sum of the  $\Delta H$  values of the constituent reactions enthalpy change  $\Delta H$  is independent of the path that a reaction follows to move from reactants

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web this activity provides a demonstration of hess law using three reactions the solubility  $\text{NaOH}$  in water the solubility  $\text{NaOH}$  in  $\text{HCl}$  and the reaction of a solution of  $\text{HCl}$  and a solution of  $\text{NaOH}$

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web hess s law worksheet answers 1 calculate  $\Delta H$  for the reaction  $\text{C}_2\text{H}_4$  g  $\text{H}_2$  g  $\text{C}_2\text{H}_6$  g from the following data  $\text{C}_2\text{H}_4$  g  $3\text{O}_2$  g  $2\text{CO}_2$  g  $2\text{H}_2\text{O}$  l  $\Delta H$  1411 kJ  $\text{C}_2\text{H}_6$  g  $3\frac{1}{2}\text{O}_2$  g  $2\text{CO}_2$  g  $3\text{H}_2\text{O}$  l  $\Delta H$  1560 kJ

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web may 8 2015 hess law lab by maya parks partners ben seufert kelsea floyd convert joules to kJ in your final answer multiply the mass by the change in temperature and the  $c$  given trial 1 use hess s law and the following equations and  $\Delta H$  values to determine the heat of reaction for the reaction  $\text{C}_2\text{H}_4$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_3\text{O}^+$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H$  55.8 kJ as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  
 $\text{1 NaOH (s)} + \text{Na}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$   
 $\text{2 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$   
 $\text{3 NaOH (s)} + \text{H}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{H}_2\text{O (l)}$