

# RoboForm User Manual (Print Version)

Copyright Siber Systems 1999-2009.

**RoboForm Version.** The manual applies to RoboForm version 6.9.93 and later.

## Installing RoboForm

**Requirements.** RoboForm works on Windows 98, Me, NT-4, 2000, 2003, XP, Vista systems. It does not work on Mac and Linux. Internet Explorer version 5.5, 6.0 or later is required. All Service Packs and security patches must be installed.

**Start Download.** Click the [Download RoboForm](#) link and a popup window will ask "Do you want to Run or Save this file?" or smth-like this. Click **Run**.

**Agree to Install RoboForm.** Once the download finishes, another popup window will ask "Do you want to run this software? RoboForm2Go by Siber Systems". Click "**Yes**" and installation will begin.

**Language and License Agreement.** Select language in which RoboForm will talk to you. Read the License Agreement and agree to it.

**Setting Up Personal Data.** In this dialog you can change folder where RoboForm ifascards and identities are stored and where RoboForm should place its program files. You can also request installation of RoboForm Adapter for Mozilla, Firefox and Netscape.

# Roboform Manual Version 7

**N Colangelo**



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web a rate law shows how the rate of a chemical reaction depends on reactant concentration for a reaction such as  $a \rightarrow b$  products the rate law generally has the form  $\text{rate} = k [a]^n$  where  $k$  is a proportionality constant called the rate constant and  $n$  is the order of the reaction with respect to  $a$  the value of  $n$  is not related to the reaction stoichiometry and



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web the rate determining slower step gives a rate law showing second order dependence on the no 2 concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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web rate laws and reaction order the relation between the rate of a reaction and the concentrations of reactants is expressed by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2\text{N}_2\text{O}_5 \rightarrow 4\text{NO}_2 + \text{O}_2$  has been found to be directly proportional to the concentration of  $\text{N}_2\text{O}_5$  text rate  $k$  n

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide  $\text{KI}$  sodium thiosulfate  $\text{Na}_2\text{S}_2\text{O}_3$  starch and

hydrogen peroxide  $\text{H}_2\text{O}_2$  were used in order to calculate the rate of reaction  
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web oct 27 2022 the rate law  $k[\text{H}][\text{OH}]$  describes a reaction that is first order in  $\text{H}$  first order in  $\text{OH}$  and second order overall example 12 3 1 writing rate laws from reaction orders an experiment shows that the reaction of

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web rate laws sometimes called differential rate laws or rate equations are mathematical expressions that describe the relationship between the rate of a chemical reaction and the concentration of its reactants as an example consider the reaction described by the chemical equation  $\text{a} + \text{b} \rightarrow \text{products}$

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web core concepts some reactions will go fast and some will go slow the speed of the reaction is its reaction rate which is dictated by a rate law in this article we will learn about reaction rates rate laws the rate constant and the reaction order

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