Name

Date \_\_\_/\_\_\_Period \_\_

# **Chapter 5 • Periodicity Activity**

#### ATOMIC RADII

Directions: Using the given data for the sizes of atoms (atomic radii) below,

- 1. Create a line graph (use a ruler) by graphing Period 2's data on the graph below.
- In a different color, create a line graph Period 3's data on the same graph.
- Label each axis, and include a key/legend.
- 4. Answer all questions.

		2	1.3	1.4	1.5	16	17	1.8		
I	30		Atomic Radii							
I	Li 123	Ble 89	B 88	C 77	N 70	O 66	F 64	Ne		
	Na 157	Mg 1.3-6	AI 125	Si 117	P 110	S 104	C1 99	Ar		
1	- K 2003	Ca. 174	Ga 125	Gie 122	As 121	Se 117	Be 114	Kr		
Ī	Rb 216	Sr 192	In 150	Sn 140	Sb 141	Te: 13-7	133	Xe		
	Cs 235	Ba 198	T1 155	Pb 154	Bi 152	Po 153	At	Ran		

Note: Mark the vertical axis from 0 - 160 (by 20's)

Trend in Atomic Size Across the Periodic Table

1.4	He Mg	18	C	24.	o 8	- 15
Non			50	100		

Horizontal Trend Questions

1. What is the trend in atomic size as you go across the periodic table?

2. This is true because there are more \_\_\_\_\_\_ (protons / electrons / layers of electrons) pulling the electron cloud in toward the nucleus.

Vertical Trend Questions

3. Examine the sizes of the Group 1 elements (H, Li, Na, K, Rb, Cs). What is the trend in atomic size as you go down a column of the periodic table?

4. This is true because there are more \_\_\_\_\_\_ (protons / electrons / layers of electrons) making the electron cloud larger.

# **Periodicity Activity Answers**

**CL Gary** 

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