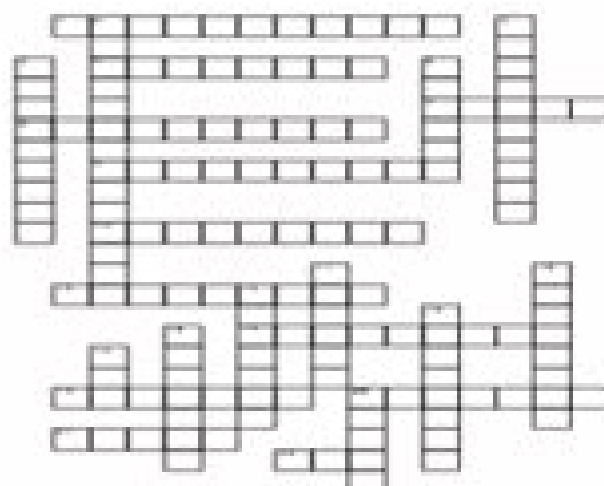


Periodic Table Assignment

Read pages 14 – 19. Complete the crossword puzzle (in pencil) and fill in the period table as instructed below.



1000

1. Synthesis elements with atomic numbers greater than 93
2. This law states that the properties of elements repeat as a function of their atomic number
3. Lanthan is in this period
4. Elements were originally placed in order of atomic ____, now they are in order of atomic ____
5. The group that consists of very unreactive gases
6. Number of elements known in 1869
7. Element 101 was named in his honour
8. The periodic table is normally drawn with 18 columns. A more concise representation has ____ columns
9. Most nonmetals are in this state
10. Nonmetal that is a liquid at SATP
11. The general name given to a column in the periodic table
12. According to figure 6, when aluminum combines with oxygen, three oxygen atoms combine with this many aluminum atoms.

1999

2. Elements that most closely follow the periodic law
3. The first scientist to organize elements based on repeating patterns
4. The modern name for alkali
6. Most elements are _____ since they fall to the left of the "staircase line"
11. The general name given to a row (left-to-right) in the periodic table
12. Iodine belongs to this group
14. The first alkali metal
15. This element was assigned a relative atomic mass of one because it is the lightest element
16. The state of matter represented by (aq)
18. Silicon is in this group
20. The staircase that separates metals and nonmetals begins below this element

1	0.0	0																	0.0	0		
2	0.0	0	0.0	0																	0.0	0
3	0.0	0	0.0	0																	0.0	0
4	0.0	0	0.0	0																	0.0	0
5	0.0	0	0.0	0																	0.0	0
6	0.0	0	0.0	0																	0.0	0
7	0.0	0	0.0	0																	0.0	0

- 1) Place a large G in boxes that house elements that are gases at S.A.T.P., and a large L in boxes that contain elements that are liquids. The rest of the elements are solids; do not write anything in these boxes.
- 2) Darken the "staircase line" that divides metals from non-metals.
- 3) Metalloids (i.e. semi-conductors) are elements that have properties midway between metals and non-metals. The metalloids are: B, Si, Ge, As, Sb, Te, Po, At. Identify the metalloids by drawing diagonal lines in these boxes.
- 4) Notice that each group (column) in the periodic table is identified at the top with a number and/or letter. There are two conventions used in the textbook. Label the groups that end in A (e.g. 1A, 2A, 3A, etc.).
- 5) Frame and colour the following groups (note: H is not an 1A metal): Alkali metals, alkaline earth metals, halogens, noble gases, transition metals, the inner transition elements (made up of lanthanides and actinides).
- 6) Create a legend for your periodic table.

Periodic Table Assignment Answer Key

JL Elias

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Periodic Table Assignment Answer Key

1. Understanding the eBook Periodic Table Assignment Answer Key
 - The Rise of Digital Reading Periodic Table Assignment Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Assignment Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Assignment Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Assignment Answer Key
 - Personalized Recommendations
 - Periodic Table Assignment Answer Key User Reviews and Ratings
 - Periodic Table Assignment Answer Key and Bestseller Lists
5. Accessing Periodic Table Assignment Answer Key Free and Paid eBooks
 - Periodic Table Assignment Answer Key Public Domain eBooks
 - Periodic Table Assignment Answer Key eBook Subscription Services
 - Periodic Table Assignment Answer Key Budget-Friendly Options

6. Navigating Periodic Table Assignment Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Assignment Answer Key Compatibility with Devices
 - Periodic Table Assignment Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Assignment Answer Key
 - Highlighting and Note-Taking Periodic Table Assignment Answer Key
 - Interactive Elements Periodic Table Assignment Answer Key
8. Staying Engaged with Periodic Table Assignment Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Assignment Answer Key
9. Balancing eBooks and Physical Books Periodic Table Assignment Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Assignment Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Assignment Answer Key
 - Setting Reading Goals Periodic Table Assignment Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Assignment Answer Key
 - Fact-Checking eBook Content of Periodic Table Assignment Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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