



# Sample Summer Learning Packet For Elementary

**LP Steffe**



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web for example take the reaction of lead ii nitrate with hydrochloric acid to form lead ii chloride and nitric acid shown below  

$$\text{Pb(NO}_3)_2 \text{ (aq)} + 2 \text{HCl (aq)} \rightarrow \text{PbCl}_2 \text{ (s)} + 2 \text{HNO}_3$$

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web you may have to consult the solubility rules write the net ionic equation for the reaction of  $\text{BaCl}_2 \text{ (aq)}$  and  $\text{Na}_2\text{SO}_4 \text{ (aq)}$  you

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web choose the correct net ionic equation representing the reaction that occurs when solutions of potassium carbonate and copper i chloride are mixed a  $\text{K}_2\text{CO}_3$  aq  $2\text{CuCl}$  aq

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web the net ionic equation for equation 1 is given as equation 3 s2 aq fe2 aq gt fes s eq 3 8 note that the phase indicator aq is optional on ionic species but is

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web cu clo4 2 is soluble all perchlorates are soluble while pbso4 is insoluble sulfates are generally soluble except for pb<sup>2</sup> ba<sup>2</sup> hg<sub>2</sub><sup>2</sup> and sr<sup>2</sup> the net ionic equation is

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web balanced chemical equation 2 naoh aq mgcl 2 aq 2 nacl aq mg oh 2 s full ionic equation 2 na aq 2 oh aq mg 2 aq 2 cl aq 2 na aq 2 cl

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web what is a titration a titration is an analytical procedure used to determine the accurate concentration of a sample by reacting it with a standard solution one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton ha b a hb strong acids e g sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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points 50 words the goal of this experiment is using the skills of precision to determine the

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web jul 19 2023 at the beginning of the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web spreadsheet computations prelab report this is the type of computations the computer does to determine first and second derivatives refer to these spreadsheet columns for the comments below pH titration graph lab data provide columns a and b

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web this problem has been solved you ll get a detailed solution from a subject matter expert that helps you learn core concepts see answer question part b acid base titration lab report i determining the concentration of sulfuric acid solution  $\text{H}_2\text{SO}_4$  by titrating it with naoh solution

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{ph} = -\log 0.100 = 1.000$  b titrant volume 12.50 ml

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web one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water in equation 1 the acid is hcl hydrochloric acid and the base is naoh sodium hydroxide when the acid and base react they form nacl sodium chloride which is also known as table salt

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