

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

$$\text{Speed of light} = (\text{frequency})(\text{wavelength}) \qquad c = f\lambda$$

$$\text{Energy} = (\text{Planck's constant})(\text{frequency}) \qquad E_{\text{photon}} = hf$$

$$\text{Energy} = \frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})} \qquad E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$\text{Total pressure of a gas} = \left(\begin{array}{c} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right) \qquad P_T = P_1 + P_2 + P_3 + \dots$$

$$(\text{Pressure})(\text{volume}) = (\text{moles})(\text{ideal gas constant})(\text{temperature}) \qquad PV = nRT$$

$$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})} \qquad \frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$(\text{Initial pressure})(\text{initial volume}) = (\text{final pressure})(\text{final volume}) \qquad P_1V_1 = P_2V_2$$

$$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})} \qquad \frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})} \qquad \frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liter of solution}} \qquad M = \frac{\text{mol}}{\text{L}}$$

$$\text{Ionization constant of water} = \left(\begin{array}{c} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left(\begin{array}{c} \text{hydroxide ion} \\ \text{concentration} \end{array} \right) \qquad K_w = [\text{H}^+][\text{OH}^-]$$

$$\left(\begin{array}{c} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left(\begin{array}{c} \text{volume of} \\ \text{solution 2} \end{array} \right) \left(\begin{array}{c} \text{molarity of} \\ \text{solution 2} \end{array} \right) \qquad V_1M_1 = V_2M_2$$

$$\text{pH} = -\log(\text{hydrogen ion concentration}) \qquad \text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

$$\text{Heat gained or lost} = (\text{mass}) \left(\begin{array}{c} \text{specific} \\ \text{heat} \end{array} \right) \left(\begin{array}{c} \text{change in} \\ \text{temperature} \end{array} \right) \qquad Q = mc_p\Delta T$$

$$\text{Enthalpy of reaction} = \left(\begin{array}{c} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left(\begin{array}{c} \text{enthalpy} \\ \text{of reactants} \end{array} \right) \qquad \Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

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Staar Eoc Chemistry Assessment Flashcard Study System Mometrix Media, 2010

SWYK on STAAR Science Gr. 8, Student Workbook Show What You Know Publishing, 2013-03-01 Assess student knowledge of the Texas Essential Knowledge and Skills TEKS for Science with two full length Assessments for each subject Questions provide students with the necessary practice needed to achieve academic success on STAAR Chapters on test taking strategies and test anxiety build students confidence and test taking skills Glossaries familiarize students with vocabulary terms and concepts found on state proficiency tests Answers are provided in the Parent Teacher Edition only

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Table of Contents Sample For Staar Testing Questions For Chemistry

1. Understanding the eBook Sample For Staar Testing Questions For Chemistry
 - The Rise of Digital Reading Sample For Staar Testing Questions For Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Sample For Staar Testing Questions For Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Sample For Staar Testing Questions For Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Sample For Staar Testing Questions For Chemistry
 - Personalized Recommendations
 - Sample For Staar Testing Questions For Chemistry User Reviews and Ratings
 - Sample For Staar Testing Questions For Chemistry and Bestseller Lists

5. Accessing Sample For Staar Testing Questions For Chemistry Free and Paid eBooks
 - Sample For Staar Testing Questions For Chemistry Public Domain eBooks
 - Sample For Staar Testing Questions For Chemistry eBook Subscription Services
 - Sample For Staar Testing Questions For Chemistry Budget-Friendly Options
6. Navigating Sample For Staar Testing Questions For Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Sample For Staar Testing Questions For Chemistry Compatibility with Devices
 - Sample For Staar Testing Questions For Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Sample For Staar Testing Questions For Chemistry
 - Highlighting and Note-Taking Sample For Staar Testing Questions For Chemistry
 - Interactive Elements Sample For Staar Testing Questions For Chemistry
8. Staying Engaged with Sample For Staar Testing Questions For Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Sample For Staar Testing Questions For Chemistry
9. Balancing eBooks and Physical Books Sample For Staar Testing Questions For Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Sample For Staar Testing Questions For Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Sample For Staar Testing Questions For Chemistry
 - Setting Reading Goals Sample For Staar Testing Questions For Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Sample For Staar Testing Questions For Chemistry
 - Fact-Checking eBook Content of Sample For Staar Testing Questions For Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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