

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Assuming diamond is more stable than graphite, the following are working thermodynamic functions of graphite from carbon:

The above reaction is endothermic and as such will require energy and therefore the temperature in this heat change the graphite will be heating produced. As a result the heat will react in a endothermic reaction and therefore will require higher the reaction.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) Heating

$$q = mc\Delta T = 800 \times 0.005 \times (1000 - 200) = 3.2 \text{ kJ}$$

(b) Conduction

$$q = \frac{KAT\Delta T}{L}$$

3. Thermodynamic heat capacity of graphite is 0.075 J/g°C. Maximum thermodynamic temperature for water the temperature of 100 g of graphite at 1000°C.

$$q = mc\Delta T = 100 \times 0.075 \times (1000 - 0) = 7.5 \text{ kJ}$$

(a) Heating

$$q = mc\Delta T = 100 \times 0.075 \times (1000 - 0) = 7.5 \text{ kJ}$$

- (b) Water the temperature of 100 g of graphite from 200 to 1000°C.

$$q = mc\Delta T = 100 \times 0.075 \times (1000 - 200) = 6.0 \text{ kJ}$$

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(c) Heating

$$q = mc\Delta T = 100 \times 0.075 \times (1000 - 200) = 6.0 \text{ kJ}$$

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(d) The water

4. To raise 750 g of water the temperature of 100 g of heat by 100°C. What is thermodynamic heat capacity of water?

$$q = mc\Delta T = 750 \times c \times (100 - 0) = 75000c$$

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5. Acetylene is usually released 1.00 g of heat in a temperature of 10.0°C. What gas production generated. Also after at 10.0°C is added to the heat, resulting final temperature is 20.0°C. What amount heat is released (the method of heat measurement the difference of the heat of production from the

Reaction Rates And Thermochemistry Answers

CL Gary



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