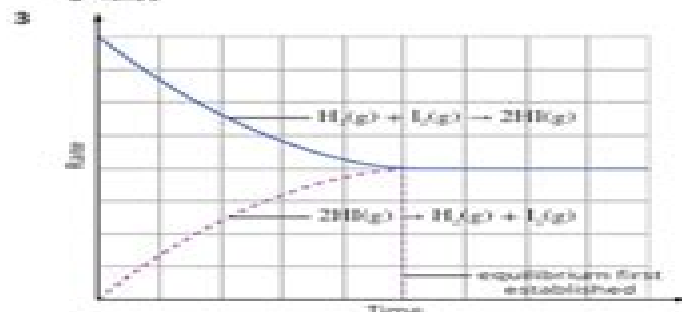


## Chapter 2 Chemical equilibrium

### 2.1 Dynamic equilibrium

#### 2.1 KEY QUESTIONS

1.
  - a An **open** system occurs when both matter and energy are exchanged between the system and the surroundings.
  - b Chemical equilibria can occur in **closed** systems.
  - c In some circumstances, a reversible reaction can form a **dynamic** equilibrium.
  - d The extent of reaction for chemical equilibria is **different** for different equilibria.
2.
  - a false
  - b false
  - c true
  - d false



- 4 closed, decreases, less, reverse, increases, equal
5.
  - a open; both energy and matter (carbon dioxide and water vapour) are exchanged with the surroundings
  - b closed; only energy can be exchanged with the surroundings. While the can is unopened, no matter is exchanged.
  - c open; both energy and matter are exchanged with the surroundings
- 6 B and F. Equilibrium can only occur in a closed system. At the molecular level, some of the sugar in the saturated sugar solution will continue to dissolve at the same rate as some is precipitated. In the unopened bottle of sparkling mineral water, the dissolved carbon dioxide will be in equilibrium with the gaseous carbon dioxide. In all of the other situations, the reverse reaction is not occurring at the same rate as the forward reaction, because the systems are open.
- 7  $130 \text{ kJ mol}^{-1}$ . The activation energy is the difference between  $390 \text{ kJ mol}^{-1}$  and  $260 \text{ kJ mol}^{-1}$ , which is  $130 \text{ kJ mol}^{-1}$ .
8.
  - a Chemical equilibrium is 'dynamic' because both forward and reverse reactions occur at the same rate. An equilibrium develops between water vapour and water when wet clothes are in a sealed bag, with water evaporating as rapidly as water vapour condenses, so the clothes remain wet.
  - b When the bag is opened, water vapour escapes and the rate of evaporation exceeds the rate of condensation. The system is not in equilibrium and the clothes dry.
9.
  - a  $0.07 \text{ mol L}^{-1}$
  - b  $0 \text{ mol L}^{-1}$
  - c  $0.03 \text{ mol L}^{-1}$
  - d  $0.08 \text{ mol L}^{-1}$
  - e 0.04 moles
- f The horizontal regions of the graph indicate that there is no change in concentration of  $\text{NO}_2$  or  $\text{N}_2\text{O}_4$ . This is when the system is at equilibrium and both the forward and reverse reactions are occurring at the same rate.
- g 6 seconds
- h As the reaction proceeds, the intensity of the brown colour will increase as the concentration of  $\text{NO}_2$  increases until equilibrium is reached. Once equilibrium is reached, the colour of the reaction mixture will remain constant.

# Reaction Rates And Equilibrium Answers Pearson Education

**Siddappa N.Byrareddy**



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### **Table of Contents Reaction Rates And Equilibrium Answers Pearson Education**

1. Understanding the eBook Reaction Rates And Equilibrium Answers Pearson Education
  - The Rise of Digital Reading Reaction Rates And Equilibrium Answers Pearson Education
  - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates And Equilibrium Answers Pearson Education
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Rates And Equilibrium Answers Pearson Education
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates And Equilibrium Answers Pearson Education
  - Personalized Recommendations
  - Reaction Rates And Equilibrium Answers Pearson Education User Reviews and Ratings
  - Reaction Rates And Equilibrium Answers Pearson Education and Bestseller Lists

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  - Reaction Rates And Equilibrium Answers Pearson Education eBook Subscription Services
  - Reaction Rates And Equilibrium Answers Pearson Education Budget-Friendly Options
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  - ePub, PDF, MOBI, and More
  - Reaction Rates And Equilibrium Answers Pearson Education Compatibility with Devices
  - Reaction Rates And Equilibrium Answers Pearson Education Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reaction Rates And Equilibrium Answers Pearson Education
  - Highlighting and Note-Taking Reaction Rates And Equilibrium Answers Pearson Education
  - Interactive Elements Reaction Rates And Equilibrium Answers Pearson Education
8. Staying Engaged with Reaction Rates And Equilibrium Answers Pearson Education
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reaction Rates And Equilibrium Answers Pearson Education
9. Balancing eBooks and Physical Books Reaction Rates And Equilibrium Answers Pearson Education
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reaction Rates And Equilibrium Answers Pearson Education
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Reaction Rates And Equilibrium Answers Pearson Education
  - Setting Reading Goals Reaction Rates And Equilibrium Answers Pearson Education
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Rates And Equilibrium Answers Pearson Education
  - Fact-Checking eBook Content of Reaction Rates And Equilibrium Answers Pearson Education
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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