

Rate Laws and Reaction Order

Rate Law → is a relation between the rate of reaction to the concentrations of the reactants

In general:



the rate law equation is:

↳ the relationship among rate, the rate constant, the initial concentration of reactants, and the orders of reaction with respect to the reactants

$$\text{Rate} = k[A]^x[B]^y[C]^z \dots$$

where k is the rate constant, i.e., it is a proportionality constant.

where x is the order of the reaction with respect to A.

y is the order of the reaction with respect to B.

z is the order of the reaction with respect to C.

↳ the exponent used to describe the relationship between the initial concentration of a particular reactant and the rate of the reaction

The sum of $x + y + z + \dots$ is called the overall order.

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