

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol/L.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol/L.

17. Explain whether the dissociation of water is an endothermic or exothermic process.
Endothermic

18. Predict, using a calculation, whether a solution of 1.0 M HCl is neutral, acidic or basic. (2)

Acidic

19. Explain by calculating pH.

20. pK_a , pH and pOH are significant notions.

Calculate pK_a of HCl .

Give your answer to two decimal places.

$$-\log(4.31 \times 10^{-4}) = 3.07$$

$$pK_a = 3.07$$

21. A solution of hydrochloric acid is mixed with water and the resulting solution is neutral at 25 °C.

Calculate the pH of the resulting solution at 65 °C.

Give your answer to two decimal places.

$$pH = 13.01$$

pH = _____ (2)

(Total: 10)

Ocr Chemistry A2 Unofficial Mark Scheme 2014

Richard Parsons



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