

Example #3 – Menus, Builtin

- Function Name has a LOV to search by name
 - Do not need to know “Function Code”
 - Use “Add Parameter” button to pass parameters for professional form functions
- Use syntax below to pass parameters to an OAF form

40 Actions for Menu1 - View Supplier Details Form

Condition Actions

Seq	Type	Description	Language	Enabled
10	Builtin	Call address book	All	☒

Action type = Builtin

Builtin Type: Launch a Function

Function Code: POS_HT_SP_B_SUPP

Function Name: Supplier: Buyer View: Supplier Search Page

Parameters:

```
=select  
OAHP=POS_HT_SP_HP&OASF=POS_HT_SP_B_SUPP&OAPS=AP_PAY_PRODUCT_BRAND')&supplier_id=')$(rem_inv_sum_folder.vendor_id value) from dual
```

Add Parameter...

Editor

```
=select  
OAHP=POS_HT_SP_HP&OASF=POS_HT_SP_B_SUPP&OAPS=AP_PAY_PRODUCT_BRAND')&supplier_id=')$(rem_inv_sum_folder.vendor_id value) from dual
```

OK Cancel Search

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Ian Pickup



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The Release 12 Primer - Shining a Light on the Release 12 World Barbara Matthews, John Stouffer, Karen Brownfield, 2008-07-18 If you've been thinking about installing or upgrading to Release 12 but need to understand more about why you should do so this is the book for you For functional users The Release 12 Primer highlights the modules that have changed the most Read about Oracle's Global Business Release Master Data Management MOAC and the Swan Interface Then see how it all comes together for the Financials and Procurement product families Supply Chain Management CRM and Projects And if you've always wanted to understand more about the underlying technology but found all the terminology too complex to sort through this book covers the technical architecture security issues and even Fusion targeting both functional and technical readers The Release 12 Primer provides the real story on the latest version of the Oracle E Business Suite and its technical underpinnings If you only buy one book on Oracle E Business Release 12 buy this one Floyd M Teter Systems Engineer JPL

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web equilibrium moles $hscn\ 2.00 \times 10^{-5}\ mol\ 3.00 \times 10^{-6}\ mol\ 1.70 \times 10^{-5}\ mol\ hscn$ knowing the number of moles of fe^{3+} and

hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species
 $\text{mol Fe}^{3+} = 1.70 \times 10^{-5} \text{ mol}$
 $\text{mol Fe}^{3+} = 8.50 \times 10^{-4} \text{ mol}$

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