

- 1. The reaction is exothermic because the products are at a lower energy state than the reactants.
- 2. The activation energy is the energy barrier that must be overcome for the reaction to proceed.
- 3. The reaction is exothermic because the products are at a lower energy state than the reactants.
- 4. The reaction is exothermic because the products are at a lower energy state than the reactants.



Figure 1



Figure 2

Explain the difference between the two graphs. The first graph shows a reaction that is exothermic, while the second graph shows a reaction that is endothermic.

Answer the questions in the margin.

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Michael Seilmaier

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