

KEY

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|----------------|-------------------|--|----------------------------|
| • Condensation | • Deposition | • Calorimeter | • Heat of reaction |
| • Freezing | • Heat | • Joule | • Le Chatelier's principle |
| • Vaporization | • Temperature | • Specific heat | • Equilibrium |
| • Melting | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$ | • Activation Energy |
| • Boiling | • Exothermic | • $q = m \cdot \Delta H_{\text{vaporization}}$ | • Catalyst |
| • Sublimation | • Endothermic | • $q = m \cdot \Delta H_{\text{fusion}}$ | |

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?

Must collide with each other, Must collide with enough energy, & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?

Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

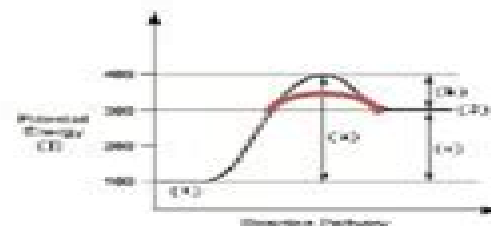


- Increase temp: *Increase, molecules collide more*
- Increase HCl: *Increase, more collisions because more reactants to collide*
- Increase pressure: *Hard to say, wouldn't affect the reactants directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call... (this one won't be on the test).*
- Remove MgCl_2 : *Increase, less products means less chance for them to interfere with reactant collisions*
- Add H_2 : *Decrease, more products mean they interfere more with reactants and the reactants would collide less.*

4. Identify the following terms with their correct letter from the graphs:

- Reactants- **I**
- Products- **2**
- Activation energy (number/ letter & amount)-
A (300 Joules)
- Is this reaction exothermic or endothermic?
Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).
Speeds it up by lowering the activation energy.



Reaction Rates And Equilibrium Test And Answers

Disha Experts



Reaction Rates And Equilibrium Test And Answers:

Solutions of Reaction Rate Equations Pertaining to Electron Irradiation of 4:1 Mixtures of N₂ and O₂

Frankline E. Niles, Edna L. Lortie, 1967 One of the Keneshea computer codes see AD 424 173 was adapted for use on the Ballistic Research Laboratories Electronic Scientific Computer Using this modified code reaction rate equations were solved for the following 15 species e NO₂ O O₂ O₃ N₂ NO O O₂ N NO N₂O NO₂ O and O₃ The calculations were made for a 4 : 1 mixture of N₂ and O₂ at 1 torr total pressure and 300K Rate constants as given by Keneshea and Fowler see AD 646 975 were used The solutions are presented as number densities versus time after the start of the irradiating electron beam A description of the modified code is presented Author **JEE Main 2020 January & September Attempt Solved Papers**

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Table of Contents Reaction Rates And Equilibrium Test And Answers

1. Understanding the eBook Reaction Rates And Equilibrium Test And Answers
 - The Rise of Digital Reading Reaction Rates And Equilibrium Test And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates And Equilibrium Test And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates And Equilibrium Test And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates And Equilibrium Test And Answers

- Personalized Recommendations
 - Reaction Rates And Equilibrium Test And Answers User Reviews and Ratings
 - Reaction Rates And Equilibrium Test And Answers and Bestseller Lists
5. Accessing Reaction Rates And Equilibrium Test And Answers Free and Paid eBooks
 - Reaction Rates And Equilibrium Test And Answers Public Domain eBooks
 - Reaction Rates And Equilibrium Test And Answers eBook Subscription Services
 - Reaction Rates And Equilibrium Test And Answers Budget-Friendly Options
 6. Navigating Reaction Rates And Equilibrium Test And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rates And Equilibrium Test And Answers Compatibility with Devices
 - Reaction Rates And Equilibrium Test And Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rates And Equilibrium Test And Answers
 - Highlighting and Note-Taking Reaction Rates And Equilibrium Test And Answers
 - Interactive Elements Reaction Rates And Equilibrium Test And Answers
 8. Staying Engaged with Reaction Rates And Equilibrium Test And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rates And Equilibrium Test And Answers
 9. Balancing eBooks and Physical Books Reaction Rates And Equilibrium Test And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates And Equilibrium Test And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Reaction Rates And Equilibrium Test And Answers
 - Setting Reading Goals Reaction Rates And Equilibrium Test And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Reaction Rates And Equilibrium Test And Answers

- Fact-Checking eBook Content of Reaction Rates And Equilibrium Test And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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