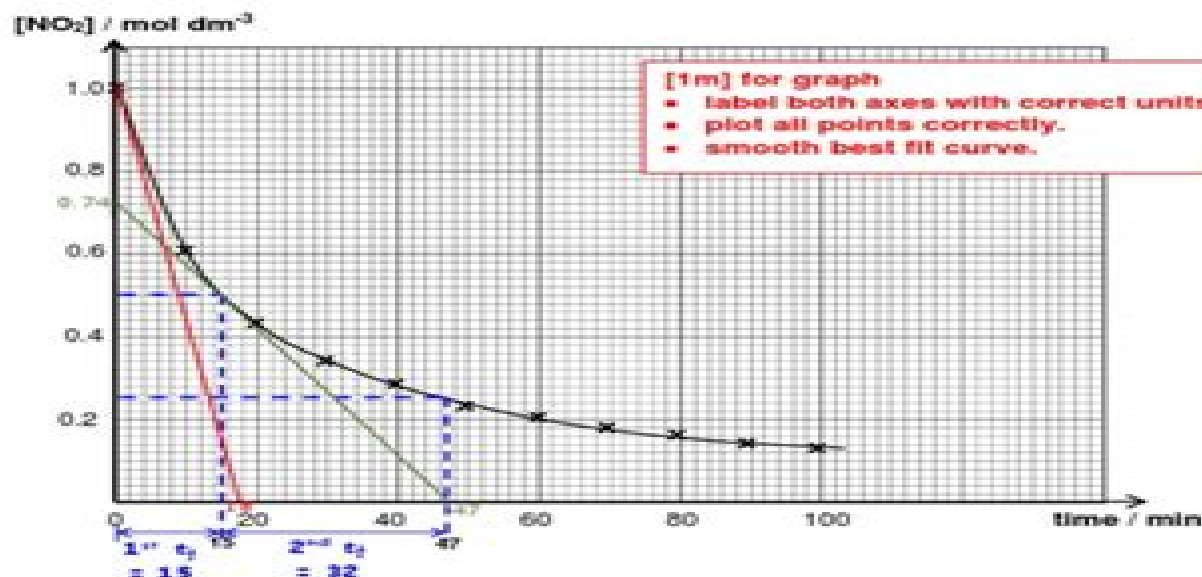


Topic 6: Reaction Kinetics (Answers)

1 (a)



[1m] Indicate and label values of two half lives clearly on the graph using construction lines

Since half-lives are not constant, the reaction is not first order. [1m]

(b) (i) Initial rate = $\frac{1.0}{18} = 0.0556 \text{ mol dm}^{-3} \text{ min}^{-1}$ accept range of 0.0500 to 0.0625

[1m] tangent line drawn at $t = 0$ + value in 3sf + units

(ii) rate when [NO₂] = 0.5 mol dm⁻³ = $\frac{0.74}{47} = 0.0157 \text{ mol dm}^{-3} \text{ min}^{-1}$

[1m] tangent line drawn at [NO₂] = 0.5 mol dm⁻³ + value in 3sf + units

(c)

[NO ₂] / mol dm ⁻³	rate / mol dm ⁻³ min ⁻¹
1.00	0.0556
0.50	0.0145

When [NO₂] increases to 2 times its original value, the rate increases to ~ 4 times its original value. } [1m]

Therefore, the order of reaction with respect to [NO₂] is 2. [1m]

Reaction Energy Reaction Kinetics Answers

Michael Brown



Reaction Energy Reaction Kinetics Answers:

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