

Answers

$Q = mc\Delta T$, where Q = heat energy, m = mass, and ΔT = change in temp.
Remember, $\Delta T = (T_{\text{final}} - T_{\text{initial}})$. Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{1086.75}{15.75(175-25)} = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = mC(T_f - T_i) = 10.0\text{g} (0.90\text{J/g}^\circ\text{C})(55-22) = 297 \text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{67500 \text{ J}}{(1500 \text{ g})(57-32)} = 1.8 \text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mC(T_f - T_i) = 100\text{g}(4.184\text{J/g}^\circ\text{C})(37 - 4) = 14000 \text{ J}$$

5. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{455 \text{ J}}{(25\text{g})(155-25)} = 0.14 \text{ J/g}^\circ\text{C}$$



Physical Science If8767 Answer Key Heat Calculations

D. V. Widder



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Physical Science Heat Energy Prentice-Hall Staff,1993 Calculations in Heat, Light and Sound ... S. M. Niven,1964

The Heat Equation D. V. Widder,1976-01-22 *The Heat Equation* *Physics II Exam File* Arthur Boyd Western,1986

The One-Dimensional Heat Equation John Rozier Cannon,2014-02-19 This is a version of Gevrey s classical treatise on the heat equations Included in this volume are discussions of initial and or boundary value problems numerical methods free boundary problems and parameter determination problems The material is presented as a monograph and or information source book After the first six chapters of standard classical material each chapter is written as a self contained unit except for an occasional reference to elementary definitions theorems and lemmas in previous chapters **New Calculations of the Specific Heat, Enthalpy and Entropy of Air** E. Justi,1937

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Table of Contents Physical Science If8767 Answer Key Heat Calculations

1. Understanding the eBook Physical Science If8767 Answer Key Heat Calculations
 - The Rise of Digital Reading Physical Science If8767 Answer Key Heat Calculations
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Answer Key Heat Calculations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science If8767 Answer Key Heat Calculations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Answer Key Heat Calculations
 - Personalized Recommendations
 - Physical Science If8767 Answer Key Heat Calculations User Reviews and Ratings
 - Physical Science If8767 Answer Key Heat Calculations and Bestseller Lists
5. Accessing Physical Science If8767 Answer Key Heat Calculations Free and Paid eBooks

- Physical Science If8767 Answer Key Heat Calculations Public Domain eBooks
 - Physical Science If8767 Answer Key Heat Calculations eBook Subscription Services
 - Physical Science If8767 Answer Key Heat Calculations Budget-Friendly Options
6. Navigating Physical Science If8767 Answer Key Heat Calculations eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science If8767 Answer Key Heat Calculations Compatibility with Devices
 - Physical Science If8767 Answer Key Heat Calculations Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science If8767 Answer Key Heat Calculations
 - Highlighting and Note-Taking Physical Science If8767 Answer Key Heat Calculations
 - Interactive Elements Physical Science If8767 Answer Key Heat Calculations
 8. Staying Engaged with Physical Science If8767 Answer Key Heat Calculations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science If8767 Answer Key Heat Calculations
 9. Balancing eBooks and Physical Books Physical Science If8767 Answer Key Heat Calculations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science If8767 Answer Key Heat Calculations
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science If8767 Answer Key Heat Calculations
 - Setting Reading Goals Physical Science If8767 Answer Key Heat Calculations
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science If8767 Answer Key Heat Calculations
 - Fact-Checking eBook Content of Physical Science If8767 Answer Key Heat Calculations
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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