AP Questions: Thermochemistry

1954 B

Substance	Standard Heat of Formation, A Hy ² in k1 mol ⁴	
C(a)	0.00	
CO ₂ (g)	-393.5	
44 st g.)	0.00	
HI ₂ O(C)	-285 85	
$O_{n(g)}$	0.00	
CallyCooking	7	

The enthalpy change for the combustion of butyric acid at 25° C, ΔH_{conb} , is -2,183.5 killojoules per mole. The combustion reaction is $C_{c}H_{c}COOH(l) + 5.0 d(e) \rightarrow 4.00 d(e) + 4.0 H_{c}O(l)$

- (a) From the above data, calculate the standard heat of formation, ΔH, for butyric acid.
- (b) Write a correctly balanced equation for the formation of butyric acid from its elements.

1988 B

	Enthalpy of	
	Combustion, A.W.	
Substance	(Ikilodos le vievoli)	
Cid	-393.5	
14(4)	-2 8.5 B	
Cally (OHI) (2)	-1366.7	
H ₂ O(0)	22	

- (a) Write a separate, balanced chemical equation for the combustion of each of the following: C(x), H₂(g), and C₁H₂OH(l). Consider the only products to be CO₂ and/or H₂O(l).
- (b) In principle, exhanol can be prepared by the following reaction:

$$2 C(x) + 2 H_2(x) + H_2O(t) \rightarrow C_0H_2OH(t)$$

Calculate the standard enthalpy change, AH, for the preparation of ethanol, as shown in the reaction above.

1988 D

An experiment is to be performed to determine the standard motar enthalpy of neutralization of a strong acid by a strong base.

Standard school laboratory equipment and a supply of standardized 1.00 molar HCl and standardized 1.00 molar NaOH are available.

- (a) What equipment would be needed?
- (b) What measurements should be taken?
- (c) Without performing calculations, describe how the resulting data should be used to obtain the standard molar enthalpy of neutralization.
- (d) When a class of students performed this experiment, the average of the results was -55.0 kilojoules per mole. The accepted value for the standard molar enthalpy of neutralization of a strong acid by a strong base is -57.7 kilojoules per mole. Propose two likely sources of experimental error that could account for the result obtained by the class.

1995 B

Propane, CaHa, is a hydrocarbon that is commonly used as fuel for cooking.

- (a) Write a balanced equation for the complete combustion of propane gas, which yields CO₂(g) and H₂O(l).
- (b) Calculate the volume of air at 30°C and 1.00 atmosphere that is needed to burn completely 10.0 grams of propane. Assume that air is 21.0 percent O₂ by volume.
- (c) The heat of combustion of propage is -2,220.1 kB/mot. Calculate the heat of formation_ΔHf, of propage given that ,ΔHf of H₂O(t) = -295.3 kB/mot and ,ΔHf of CO₂(g) = -393.5 kB/mot.
- (d) Assuming that all of the heat evolved in burning 30.0 grams of propane is transferred to 8.00 killograms of water (specific heat = 4.18 J/g·K), calculate the increase in temperature of water.

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