

12

THE BEHAVIOR OF GASES

Conceptual Curriculum

- Concrete concepts
- More abstract concepts or math/problem-solving

Standard Curriculum

- Core content
- Extension topics

Honors Curriculum

- ◆ Core honors content
- ◇ Options to accelerate

SECTION 12.1 THE PROPERTIES OF GASES (pages 327–328)



This section describes the properties of gas particles and explains how the kinetic energy of gas particles relates to Kelvin temperature.

► Kinetic Theory Revisited (pages 327–328)

- What theory explains the behavior of gases with respect to conditions such as temperature and pressure? kinetic theory
- If you notice that a sealed bag of potato chips bulges when placed near a sunny window, what can you hypothesize about the relationship between the temperature and pressure of an enclosed gas?
An increase in temperature causes an increase in the pressure exerted by the gas.
- List three basic assumptions of the kinetic theory about the properties of gases.
 - Gas particles are so small in relation to the distances between them that their individual volumes can be assumed to be insignificant.
 - No attractive or repulsive forces exist between gas particles.
 - Gas particles are in constant random motion, traveling in independent straight paths until they collide with each other or other objects.
- Circle the letter next to each sentence that is true concerning the compressibility of gases.
 - ☒ The large relative distances between gas particles means that there is considerable empty space between them.
 - ☒ The assumption that gas particles are far apart explains gas compressibility.
 - ☒ Compressibility is a measure of how much the volume of matter decreases under pressure.
 - ☐ Energy is released by a gas when it is compressed.
- Look at Figure 12.1 on page 327. How does an automobile air bag protect the crash dummy from being broken as a result of impact?
The gases used to inflate the airbag are compressible and are able to absorb a considerable amount of energy.

Section 3 Behavior Of Gases Answer Key

Emilie Sanchez



Section 3 Behavior Of Gases Answer Key:

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The Expanded Social Scientist's Bestiary

Denis Charles Phillips, 2000 *The Expanded Social Scientist's Bestiary* addresses a number of important theoretical and philosophical issues in the social sciences from the perspective of contemporary philosophy of science. The book discusses and critiques the various arguments that purport to establish that it is a mistake to believe that a naturalistic social science is a social science that in some way resembles the natural sciences can be produced. It is intended to guide social scientists, researchers, teachers, and students so that they will not fall victim to the beasts they will encounter in the course of their inquiries. Such beasts include holism, post-positivistic work in the philosophy of science, Kuhnian relativism, the denial of objectivity, and value neutrality, hermeneutics, and several others, both good and bad. This expanded and revised edition contains four new chapters tackling such contemporary beasts as Popperian rules, narrative research, and various forms of constructivism. The chapters presented in this volume are as far as possible self-contained so that each chapter can be consulted without the necessity of having read the others, thus making this volume an invaluable guide for faculty members and graduate students in the whole of the social sciences and related applied fields.

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Study Guide for Chemical Principles Thomas Elliott Taylor, 1979

Fundamentals of Fire Protection Arthur Cote, 2011-02-28 Up-to-date, broad-based training for fire service candidates and in-service professionals. Comprehensive coverage from fire basics to fire department operations and based on objectives established by the National Fire Academy. Written by experienced fire service faculty from colleges and fire departments. *Fundamentals of Fire Protection* provides a solid introduction to the full range of fire protection topics. Designed for classroom instruction or self-study, this authoritative resource is a suggested text for the model FESHE curriculum course.

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